

Markscheme

May 2025

Chemistry

Higher level

Paper 2

© International Baccalaureate Organization 2025

All rights reserved. No part of this product may be reproduced in any form or by any electronic or mechanical means, including information storage and retrieval systems, without the prior written permission from the IB. Additionally, the license tied with this product prohibits use of any selected files or extracts from this product. Use by third parties, including but not limited to publishers, private teachers, tutoring or study services, preparatory schools, vendors operating curriculum mapping services or teacher resource digital platforms and app developers, whether fee-covered or not, is prohibited and is a criminal offense.

More information on how to request written permission in the form of a license can be obtained from <https://ibo.org/become-an-ib-school/ib-publishing/licensing/applying-for-a-license/>.

© Organisation du Baccalauréat International 2025

Tous droits réservés. Aucune partie de ce produit ne peut être reproduite sous quelque forme ni par quelque moyen que ce soit, électronique ou mécanique, y compris des systèmes de stockage et de récupération d'informations, sans l'autorisation écrite préalable de l'IB. De plus, la licence associée à ce produit interdit toute utilisation de tout fichier ou extrait sélectionné dans ce produit. L'utilisation par des tiers, y compris, sans toutefois s'y limiter, des éditeurs, des professeurs particuliers, des services de tutorat ou d'aide aux études, des établissements de préparation à l'enseignement supérieur, des fournisseurs de services de planification des programmes d'études, des gestionnaires de plateformes pédagogiques en ligne, et des développeurs d'applications, moyennant paiement ou non, est interdite et constitue une infraction pénale.

Pour plus d'informations sur la procédure à suivre pour obtenir une autorisation écrite sous la forme d'une licence, rendez-vous à l'adresse <https://ibo.org/become-an-ib-school/ib-publishing/licensing/applying-for-a-license/>.

© Organización del Bachillerato Internacional, 2025

Todos los derechos reservados. No se podrá reproducir ninguna parte de este producto de ninguna forma ni por ningún medio electrónico o mecánico, incluidos los sistemas de almacenamiento y recuperación de información, sin la previa autorización por escrito del IB. Además, la licencia vinculada a este producto prohíbe el uso de todo archivo o fragmento seleccionado de este producto. El uso por parte de terceros —lo que incluye, a título enunciativo, editoriales, profesores particulares, servicios de apoyo académico o ayuda para el estudio, colegios preparatorios, desarrolladores de aplicaciones y entidades que presten servicios de planificación curricular u ofrezcan recursos para docentes mediante plataformas digitales—, ya sea incluido en tasas o no, está prohibido y constituye un delito.

En este enlace encontrará más información sobre cómo solicitar una autorización por escrito en forma de licencia: <https://ibo.org/become-an-ib-school/ib-publishing/licensing/applying-for-a-license/>.

Subject Details: Chemistry Higher Level Paper 2 Markscheme

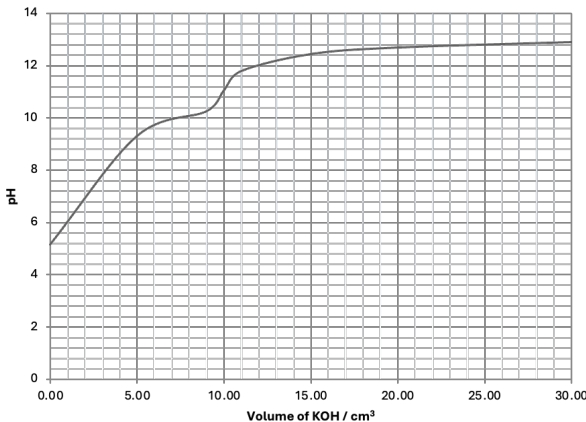
Candidates are required to answer **ALL** questions. Maximum total = **[90 marks]**.

1. Each row in the “Question” column relates to the smallest subpart of the question.
2. The maximum mark for each question subpart is indicated in the “Total” column.
3. Each marking point in the “Answers” column is shown by means of a tick (✓) at the end of the marking point.
4. A question subpart may have more marking points than the total allows. This will be indicated by “**max**” written after the mark in the “Total” column. The related rubric, if necessary, will be outlined in the “Notes” column.
5. An alternative word is indicated in the “Answers” column by a slash (/). Either word can be accepted.
6. An alternative answer is indicated in the “Answers” column by “**OR**”. Either answer can be accepted.
7. An alternative markscheme is indicated in the “Answers” column under heading **ALTERNATIVE 1** etc. Either alternative can be accepted.
8. Words inside chevrons « » in the “Answers” column are not necessary to gain the mark.
9. Words that are underlined are essential for the mark.
10. The order of marking points does not have to be as in the “Answers” column, unless stated otherwise in the “Notes” column.
11. If the candidate’s answer has the same “meaning” or can be clearly interpreted as being of equivalent significance, detail and validity as that in the “Answers” column then award the mark. Where this point is considered to be particularly relevant in a question it is emphasized by **OWTTE** (or words to that effect) in the “Notes” column.
12. Remember that many candidates are writing in a second language. Effective communication is more important than grammatical accuracy.
13. Occasionally, a part of a question may require an answer that is required for subsequent marking points. If an error is made in the first marking point then it should be penalized. However, if the incorrect answer is used correctly in subsequent marking points then **follow through** marks should be awarded. When marking, indicate this by adding **ECF** (error carried forward) on the script.
14. Do **not** penalize candidates for errors in units or significant figures, **unless** it is specifically referred to in the “Notes” column.
15. If a question specifically asks for the name of a substance, do not award a mark for a correct formula unless directed otherwise in the “Notes” column. Similarly, if the formula is specifically asked for, do not award a mark for a correct name unless directed otherwise in the “Notes” column.
16. If a question asks for an equation for a reaction, a balanced symbol equation is usually expected, do not award a mark for a word equation or an unbalanced equation unless directed otherwise in the “Notes” column.
17. Ignore missing or incorrect state symbols in an equation unless directed otherwise in the “Notes” column.

Question			Answers	Notes	Total
1	(a)	(i)	✓	Accept any combination of dots or crosses to represent electrons, or lines to represent electron pairs.	1
1	(a)	(ii)	Hybridization: sp ✓ Sigma bonds: 2 AND Pi bonds: 2 ✓		2
1	(a)	(iii)	Molecular geometry: linear ✓ Explanation: there are two electron domains around the carbon/central atom OR 2 bonding pairs and no lone pairs of electrons around the central atom «these take up the orientation where there is least electrostatic repulsion» ✓		2
1	(a)	(iv)	Partial positive charge: hydrogen/H AND Partial negative charge: nitrogen/N ✓	Accept carbon/C for “Partial positive charge”.	1

Question			Answers	Notes	Total
1	(a)	(v)	N₂ «only» has London/dispersion forces AND HCN has dipole-dipole «and London(dispersion) »forces OR HCN has dipole-dipole force AND N₂ does not ✓ intermolecular forces are weaker in N₂ OR dipole-dipole forces stronger «than London(dispersion) forces»✓	<i>Accept instantaneous / transient / induced dipole attractions for London forces</i> <i>Do not accept van der Waals' forces for M1.</i>	2
1	(b)	(i)	HCN(aq) $\rightleftharpoons$ H⁺(aq) + CN⁻(aq) ✓ OR HCN(aq) + H₂O(l) $\rightleftharpoons$ H₃O⁺(aq) + CN⁻(aq) ✓		1

Question			Answers	Notes	Total
1	(b)	(ii)	Any two of: measure pH «HCN has» higher pH ✓ measure electrical conductivity «HCN has» lower electrical conductivity ✓ react with a metal/metal oxide/metal hydroxide/metal carbonate/metal hydrogencarbonate «HCN» reacts more slowly ✓	Award [1] for two correct methods without an observation. Accept neutralization/titration with a strong base; temperature rise is lower/equivalence point > 7 for HCN Accept it would turn universal indicator yellow/orange rather than red. Accept other indicators with appropriate colour changes. Do not accept answers that refer to amounts rather than rates.	2 Max
1	b	(iii)	$[H^+] = \langle \sqrt{0.100 \times 4.90 \times 10^{-10}} \rangle$ = 7.00×10^{-6} «mol dm⁻³» ✓ pH «= -log(7.00 × 10⁻⁶)» = 5.15 ✓	Award [2] for correct final answer.	2

Question			Answers	Notes	Total
1	(b)	(iv)	 sigmoidal curve AND equivalence point > 7 ✓ curve starts at $4 < \text{pH} < 6$ AND end at $11 < \text{pH} < 14$ ✓ vertical rise at volume of 10 cm^3 ✓		3
1	(b)	(v)	equivalence point on student graph labelled “Salt” ✓		1

Question			Answers	Notes	Total
1	(b)	(vi)	amount neutralised «= $c \times V = 0.100 \times 0.0200$ » = 0.00200 «mol» ✓ $\Delta T \left\langle \left(= \frac{0.00200 \times 13300}{50.0 \times 4.18} \right) \right\rangle = 0.127$ «K / °C» ✓	Award [2] for correct final answer.	2
1	(c)	(i)	coordination ✓	Accept dative/coordination covalent.	1
1	(c)	(ii)	+2 ✓	Accept II for +2 but do not accept 2+ or 2.	1
1	(c)	(iii)	Any 3 of: partially filled d-orbitals ✓ «ligands cause the energies of the» d-orbitals to split ✓ electrons can absorb light energy as they move from lower to upper level / are promoted ✓ wavelength /energy gap corresponds to visible region ✓	Do not award final marking point for colour observed is complementary colour of light absorbed.	3 Max

Question			Answers	Notes	Total
1	(c)	(iv)	$\frac{[Fe(CN)_6]^{4-}}{[Fe^{2+}] \cdot [CN^-]^6} \quad \checkmark$ $[CN^-] = 6 \times [Fe^{2+}] = 1.32 \times 10^{-6} \text{ «mol dm}^{-3}\text{»} \quad \checkmark$ $K = \left\langle \frac{[Fe(CN)_6]^{4-}}{[Fe^{2+}] \cdot [CN^-]^6} = \frac{[1.00]}{[2.2 \times 10^{-7}] \cdot [1.32 \times 10^{-6}]^6} = \frac{[1.00]}{[2.2 \times 10^{-7}] \cdot [5.29 \times 10^{-36}]} = \frac{[1.00]}{[1.16 \times 10^{-42}]} \right\rangle = 8.59 \times 10^{41} \quad \checkmark$	Award [3] for correct final answer.	3
1	(c)	(v)	$E^\ominus$ less positive/lower than «+0.77 V for» Fe^{3+}/Fe^{2+} $\checkmark$ the less positive $E^\ominus$ the weaker the oxidizing agent / stronger the reducing agent OR thus, a weaker oxidizing agent $\checkmark$		2
1	(d)	(i)	$\begin{array}{c} H \\ \\ H-C-C \equiv N \\ \\ H \end{array} \quad / \quad H_3C-C \equiv N \quad \checkmark$		1
1	(d)	(ii)	S_N2 AND primary halogenoalkane $\checkmark$	Accept “forms a primary carbocation”.	1

Question			Answers	Notes	Total
1	(d)	(iii)	«carry out the reaction with» different concentrations of the nucleophile/CN^- «with all other conditions unchanged» ✓ rate varies with $[\text{nucleophile}]/[\text{CN}^-]$ then $\text{S}_{\text{N}}2$ OR rate does not vary with $[\text{nucleophile}]/[\text{CN}^-]$ then $\text{S}_{\text{N}}1$ ✓		2
2	(a)	(i)	$^{203}_{81}\text{Tl}$ ✓		1
2	(a)	(ii)	«A_{r}» = $(0.3 \times 203) + (0.7 \times 205)$ ✓ «A_{r}» = 204.40 ✓	Award [2] for correct final answer. Do not accept 204.38 (value in data booklet) for M2.	2

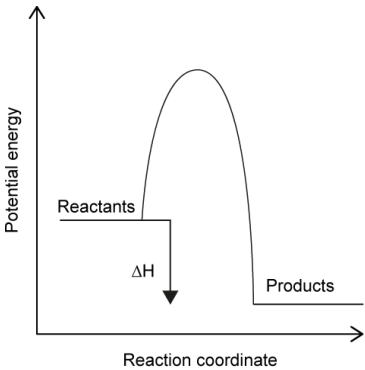
Question			Answers	Notes	Total
2	(b)	(i)	<i>Covalent bond between: sulfur/S AND oxygen/O ✓</i> <i>Ionic bond between: Tl⁺ AND SO₄²⁻ ✓</i>		2
2	(b)	(ii)	covalent bonding has shared electrons « between atoms» AND ionic bonding involves electron transfer «between atoms» ✓	<i>Accept “gain and loss”/donate instead of “transfer”.</i>	1
2	(b)	(iii)	lattice enthalpy ✓		1
2	(b)	(iv)	$2 \text{ TlOH(s)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow \text{Tl}_2\text{SO}_4\text{(aq)} + 2 \text{ H}_2\text{O(l)}$ correct formulas of products ✓ correct equation ✓		2

Question			Answers	Notes	Total
2	(b)	(v)	amount TlOH $\langle \langle = \frac{10.0}{221.39} \rangle \rangle = 0.0452 \checkmark$ amount $\text{H}_2\text{SO}_4 = \langle 0.0452 \times \frac{1}{2} \rangle = 0.0226 \checkmark$ volume of $\text{H}_2\text{SO}_4 \langle \langle = \frac{0.0226}{2.00} \rangle \rangle$ $= 0.0113 \text{ dm}^3 / 11.3 \text{ cm}^3 \checkmark$	Award [3] for correct final answer.	3
2	(b)	(vi)	yes AND in same group as aluminium/group 13 «which has an amphoteric oxide» OR no AND metallic character increases going down a group «so will have basic oxide like most metals» $\checkmark$		1

Question			Answers	Notes	Total
2	(b)	(vii)	ALTERNATIVE 1: place thallium in a solution of copper sulfate ✓ if reaction occurs then thallium is more reactive OR if no reaction occurs then copper is more reactive ✓ ALTERNATIVE 2: place copper in a solution of thallium sulfate ✓ if reaction occurs then copper is more reactive OR if no reaction occurs then thallium is more reactive ✓		2

Question			Answers	Notes	Total
2	(b)	(vii i)	<i>Anode:</i> oxygen AND water is more easily oxidized than sulfate ion ✓ <i>Cathode:</i> thallium AND thallium is more easily reduced than water ✓		2
2	(c)	(i)	limit of convergence «of continuum» ✓	<i>Accept high energy limit instead of limit of convergence.</i>	1

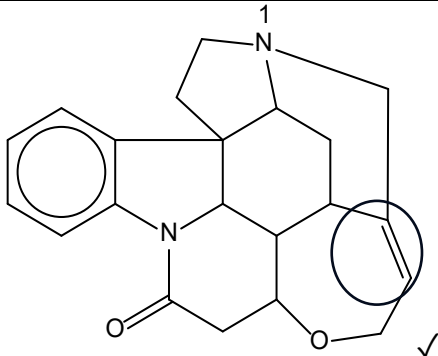
Question			Answers	Notes	Total
2	(c)	(ii)	$\text{IE} \left(\left\langle = \frac{589 \times 10^3 \text{ J mol}^{-1}}{6.02 \times 10^{23} \text{ mol}^{-1}} \right\rangle \right)$ $= 9.78 \times 10^{-19} \text{ J atom}^{-1} \checkmark$ $\text{frequency} \left(\left\langle = \frac{E}{h} = \frac{9.78 \times 10^{-19} \text{ J}}{6.63 \times 10^{-34} \text{ Js}} \right\rangle \right) =$ $1.48 \times 10^{15} \text{ s}^{-1} \checkmark$ $\text{wavelength} \left(\left\langle = \frac{c}{f} = \frac{3.00 \times 10^8 \text{ m s}^{-1}}{1.48 \times 10^{15} \text{ s}^{-1}} \right\rangle \right)$ $= 2.03 \times 10^{-7} \text{ m} / 203 \text{ nm} \checkmark$	Award [3] for correct final answer.	3
2	(c)	(iii)	nuclear charge of Tl less than that of Pb / increasing «effective» nuclear charge from Tl to Pb $\checkmark$ shielding effect/occupied energy levels stay the same $\checkmark$	Do not accept the atomic radius of lead is less than that of thallium for M1. Do not accept that the loss of an electron from thallium leaves a complete subshell.	2

Question			Answers	Notes	Total
3	(a)	(i)	$\text{Cl}_2(\text{g}) \rightarrow 2 \text{Cl}\cdot(\text{g}) \checkmark$	Accept $\frac{1}{2} \text{Cl}_2(\text{g}) \rightarrow \text{Cl}\cdot(\text{g})$. Do not accept equations without a dot indicating the radical.	1
3	(a)	(ii)	homolytic $\checkmark$		1
3	(a)	(iii)	«bond breaking» = $\text{CO} + \text{Cl}-\text{Cl} / 1077 + 242 / 1319 \text{ «kJ mol}^{-1}\text{»} \checkmark$ «bond forming» = $\text{C}=\text{O} + 2 \times \text{C}-\text{Cl} / 804 + 2 \times 324 / 1452 \text{ «kJ mol}^{-1}\text{»} \checkmark$ « $\Delta H = 1319 - 1452 =$ » $-133 \text{ «kJ mol}^{-1}\text{»} \checkmark$	Award [3] for correct final answer.	3
3	(a)	(iv)	 exothermic reaction profile $\checkmark$ ΔH correct AND labelled reactants and products $\checkmark$		2

Question			Answers	Notes	Total
3	(a)	(v)	entropy change «= 284 - 223 - 198» = -137 «J mol ⁻¹ K ⁻¹ » ✓		1
3	(a)	(vi)	$\Delta G^{\circ} \llcorner \Delta H^{\circ} - T\Delta S^{\circ} = -133000 - (298 \times -137)$ $= -133000 + 40826 \llcorner = -92174 \llcorner \text{J mol}^{-1} \llcorner$ $\ln K = \llcorner \llcorner - \frac{\Delta G^{\circ}}{R.T} = - \frac{-92174}{8.31 \times 298} \llcorner \llcorner = 37.2 \llcorner$ $K \llcorner e^{37.2} = 1 \times 10^{16} \llcorner \text{too large for equilibrium} \llcorner$	Accept answers in the range 1.2-1.5 $\times 10^{16}$.	3

Question			Answers	Notes	Total
3	(b)	(i)	change in pressure per unit time «at constant volume» OR change in volume per unit time «at constant pressure» ✓	<i>Do not accept change in pH or titrating.</i>	1
3	(b)	(ii)	incorrect orientation ✓ «collision» energy is less than activation energy ✓		2
3	(b)	(iii)	«rate =» $k [\text{Cl}_2\text{CO}] [\text{H}_2\text{O}]$ ✓		1
3	(b)	(iv)	$[\text{H}_2\text{O}] \gg [\text{Cl}_2\text{CO}] / [\text{H}_2\text{O}]$ remains constant ✓		1
3	(b)	(v)	x-axis: $1/T$ AND y-axis: $\ln k$ ✓		1
3	(b)	(vi)	gradient $\left(\frac{\ln k}{1/T} = -\frac{E_a}{R} = -\frac{12000}{8.31} \right) = -1444 \text{ «K»}$ ✓	Accept $1.4 - 1.5 \times 10^3 \text{ «K»}$.	1
3	(b)	(vii))	higher temperature curve added with maximum lower and shifted to right ✓	<i>Do not award mark unless area under curve approximately the same, the curve starts at (0,0) and it remains above the x-axis.</i>	1

Question			Answers	Notes	Total
3	(b)	(vii i)	E_a appropriately indicated ✓ explanation stating proportion of collisions/particles with $E > E_a$ OR area under curve when $E > E_a$ increases at higher temperature ✓		2
3	(b)	(ix)	not consumed in reaction / unchanged ✓ «reaction» rate increased «under same conditions» ✓		2
3	(c)	(i)	fragmentation ✓	Accept molecule unstable/disintegrates in instrument/on ionization.	1
3	(c)	(ii)	Cl-CO ⁺ ✓	Do not award the mark if the + sign is missing.	1
3	(c)	(iii)	<i>IR</i> : «absorption band due to C-Cl at » 600-800 cm ⁻¹ OR «absorption band due to C=O at » 1700-1750 cm ⁻¹ ✓ ¹ H NMR: no signals «as no H atoms in molecule» ✓		2
4	(a)	(i)	«tertiary» amino ✓	Accept amine.	1

Question			Answers	Notes	Total
4	(a)	(ii)	Direction of pH change: increase AND Reason for change: accepts hydrogen ion/proton «from water» ✓	Accept increase AND donates an electron pair. Accept increase AND acts as a base.	1
4	(a)	(iii)			1
4	(a)	(iv)	7 ✓		1
4	(a)	(v)	a «carbon» atom bonded to four different atoms «or groups of atoms» ✓	Accept a carbon atom that is not superimposable on its mirror image. Accept four different substituents/functional groups. Do not accept answers that refer to optical properties.	1

Question			Answers	Notes	Total
4	(b)		theoretical yield $\langle (= 48.73 \times \frac{2 \times 334.4 + 98.1}{2 \times 334.4} = 48.73 \times \frac{766.9}{668.8}) \rangle$ $= 55.87 \text{ g} \checkmark$ percentage yield $\langle (= 100 \times \frac{51.41}{55.87}) \rangle \langle = 100 \times 51.41 / 55.87 \rangle = 92.00 \text{ «\%» } \checkmark$	<i>Award [2] for correct final answer.</i>	2
