



Diploma Programme
Programme du diplôme
Programa del Diploma

Chemistry

Standard level

Paper 1A



16 May 2025

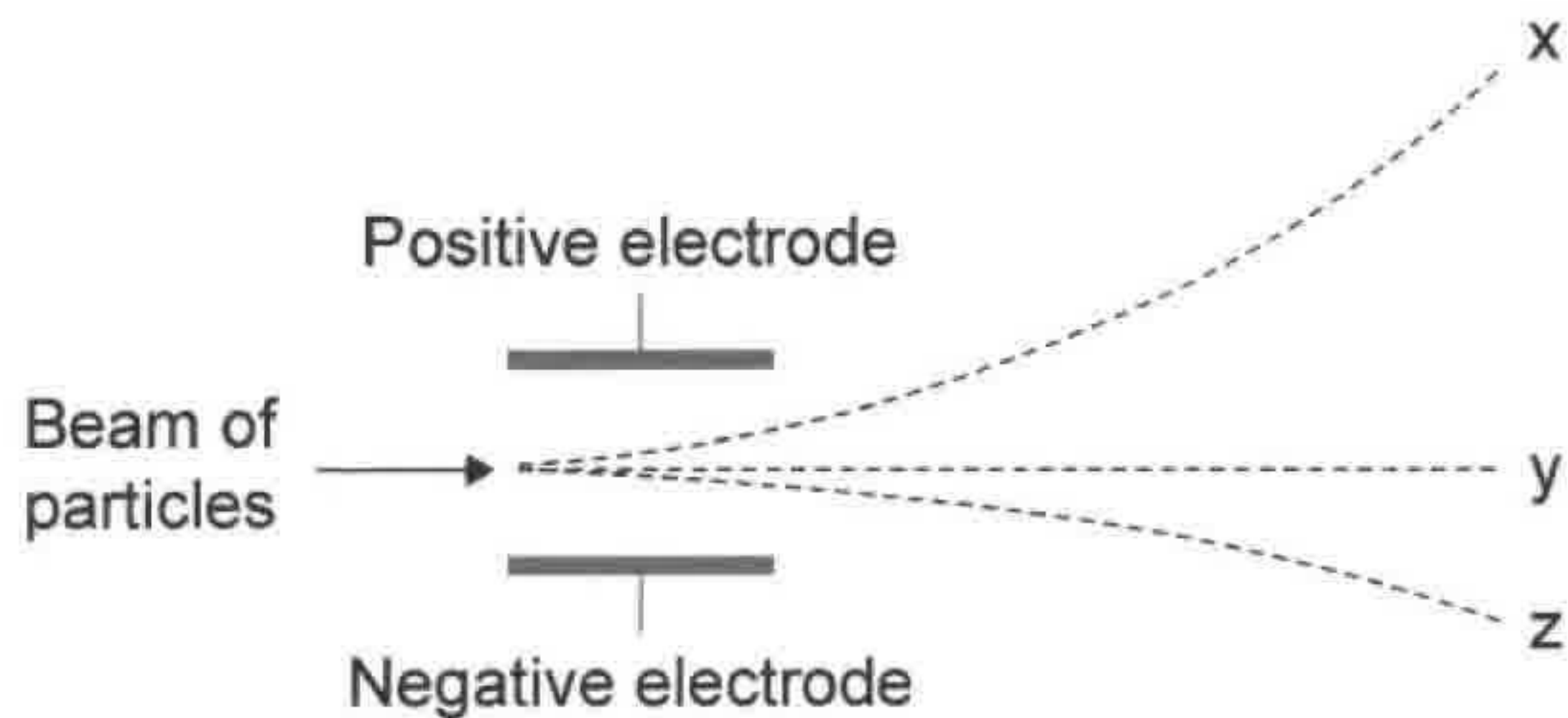
Zone A afternoon | **Zone B** afternoon | **Zone C** afternoon

1 hour 30 minutes [Paper 1A and Paper 1B]

Instructions to candidates

- Do not open this examination paper until instructed to do so.
- Answer all questions.
- For each question, choose the answer you consider to be the best and indicate your choice on the answer sheet provided.

1. The diagram below shows the behaviour of protons, neutrons and electrons in an electric field.

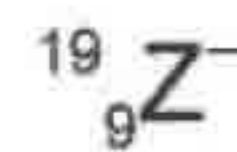
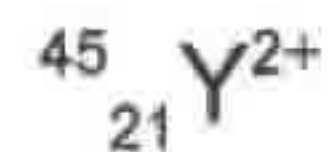
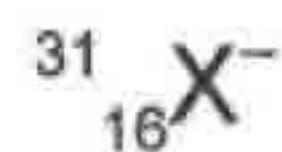
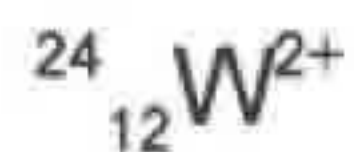


What is the correct identity of each particle?

	Protons	Electrons	Neutrons
A.	y	z	x
B.	x	y	z
C.	z	x	y
D.	x	z	y

C.	z	x	y
D.	x	z	y

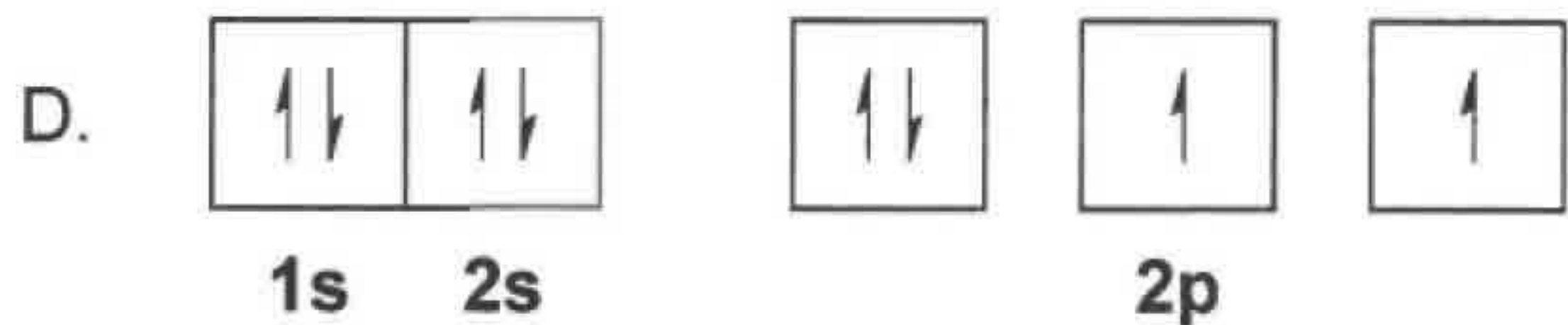
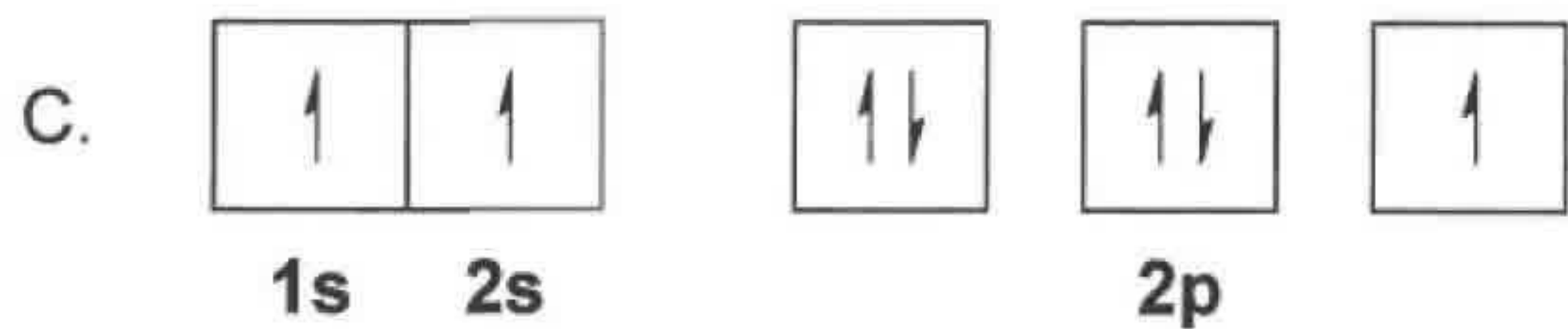
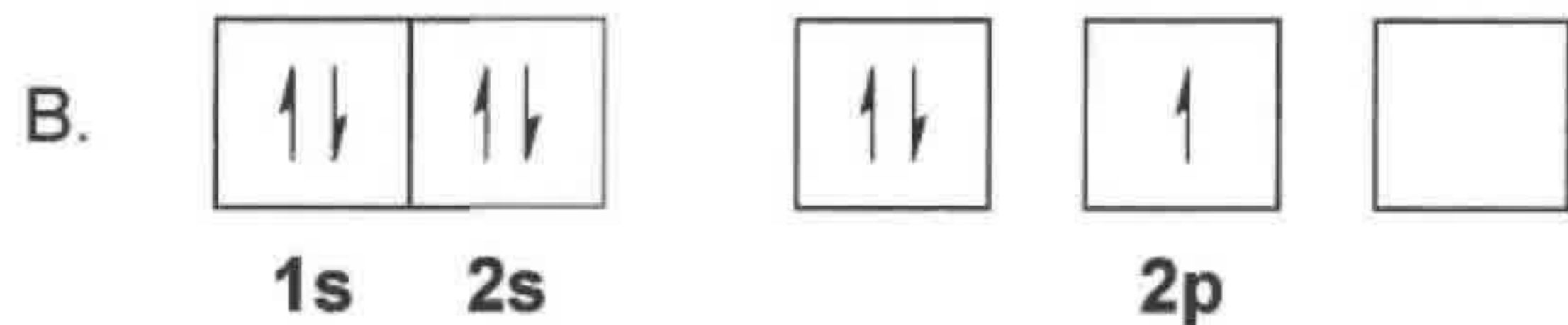
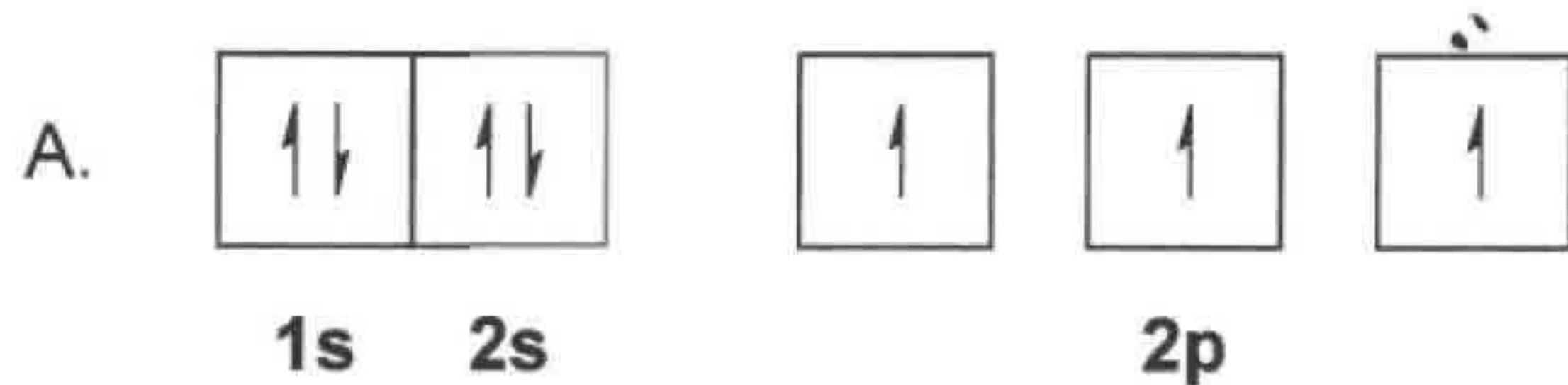
2. Which of the following species have the same number of outer electrons? The letters **do not** represent symbols of elements.



- A. X and Y
- B. X and W
- C. Z and Y
- D. Z and W



3. Which diagram shows the correct electron configuration of a nitrogen atom in the ground state?



1s**2s****2p**

4. Which row shows a gas that would deviate the least from ideal gas behaviour?

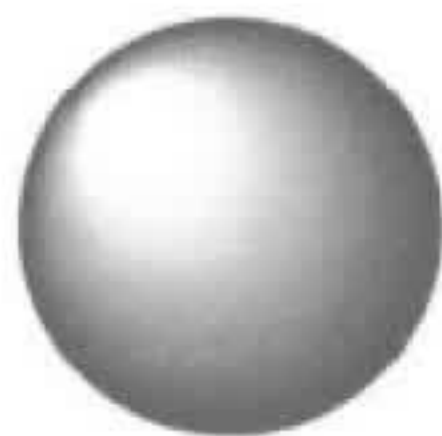
	Gas	Pressure	Temperature
A.	Phosphine, PH_3	Low	High
B.	Ammonia, NH_3	Low	High
C.	Phosphine, PH_3	High	Low
D.	Ammonia, NH_3	High	Low

5. Which row correctly shows the shape of s and p atomic orbitals?

A.



s

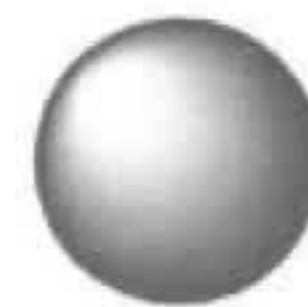


p

B.



s

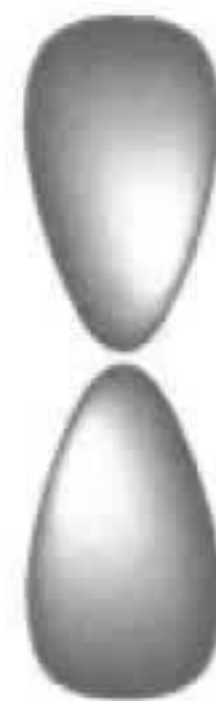


p

C.



s



p

D.



s



p





s



p

s

p

6. The volume of a container filled with a fixed amount of gas is increased by 100 % at constant temperature. How does this affect the pressure in a container? $pV = nRT$
- A. Pressure increases by 100 %.
 - B. Pressure decreases by 100 %.
 - C. Pressure increases by 50 %.
 - D. Pressure decreases by 50 %.
7. Which row shows the correct formula and electron configuration of the species in calcium nitride?

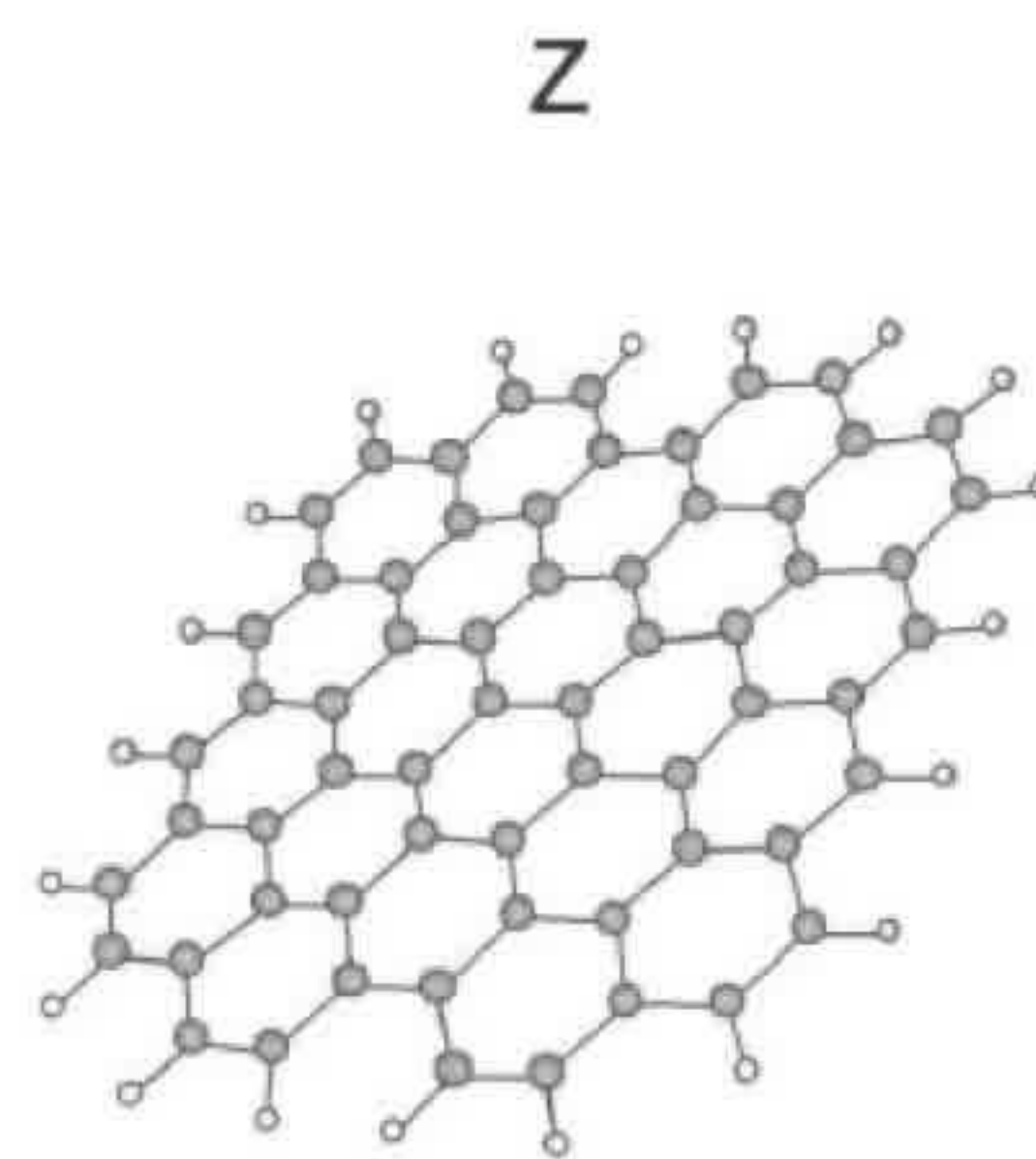
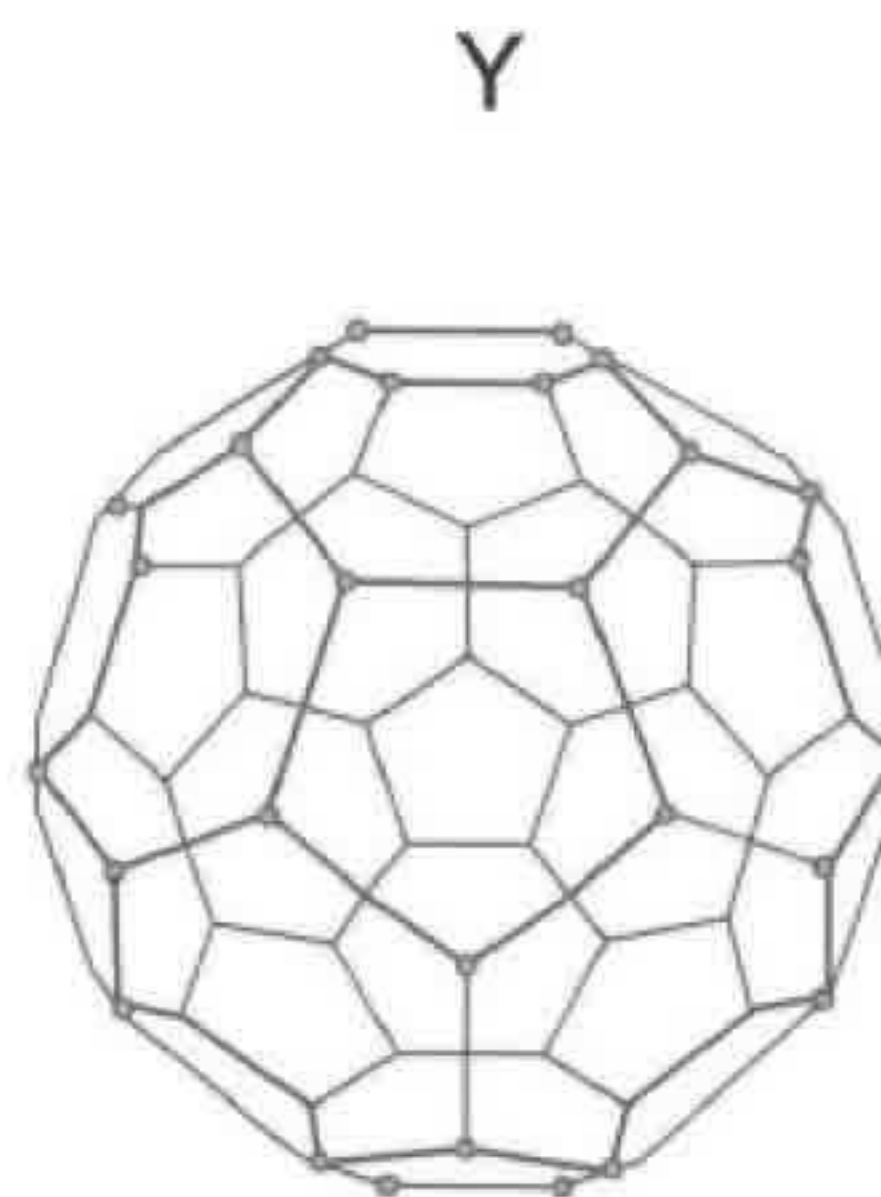
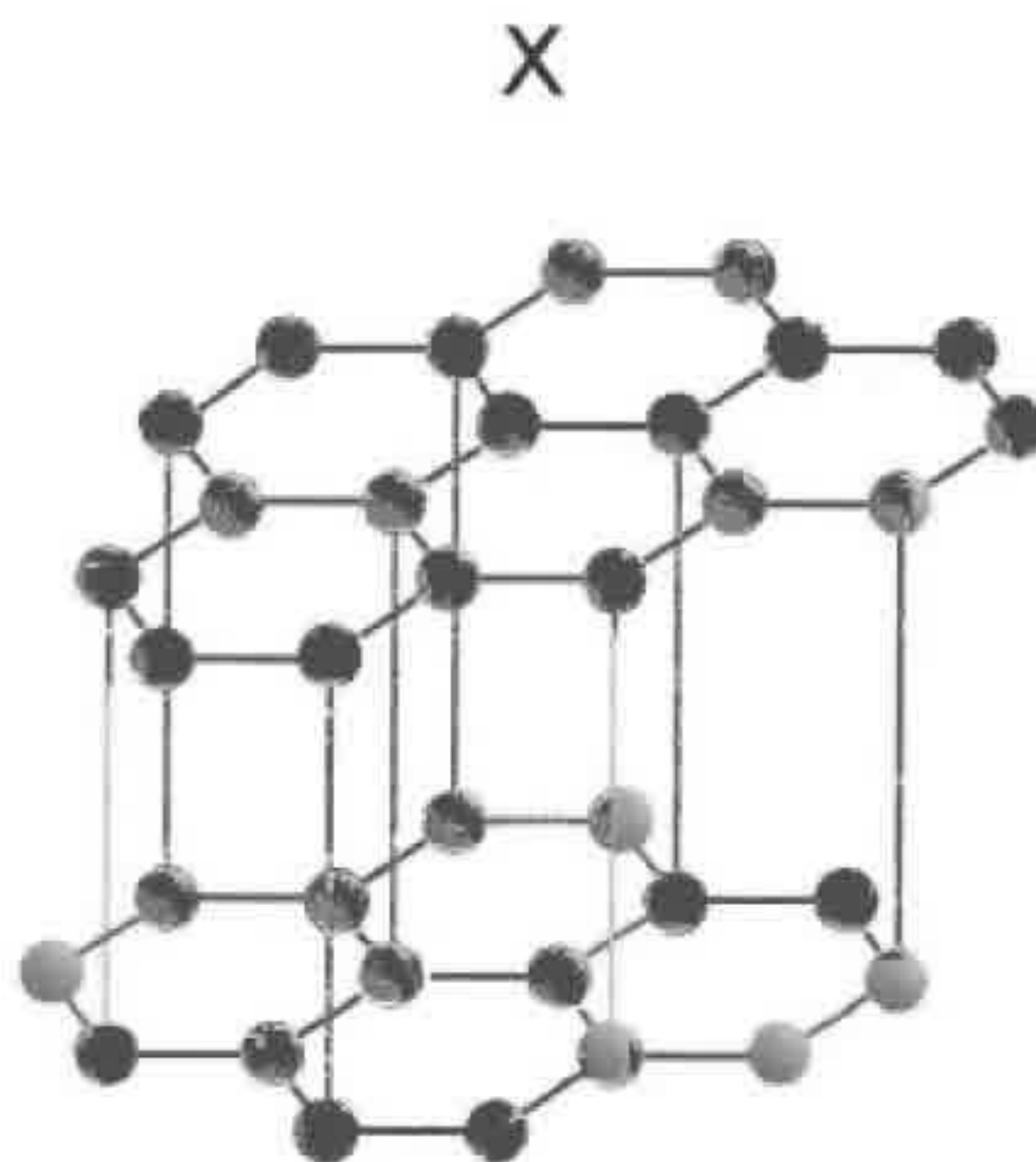
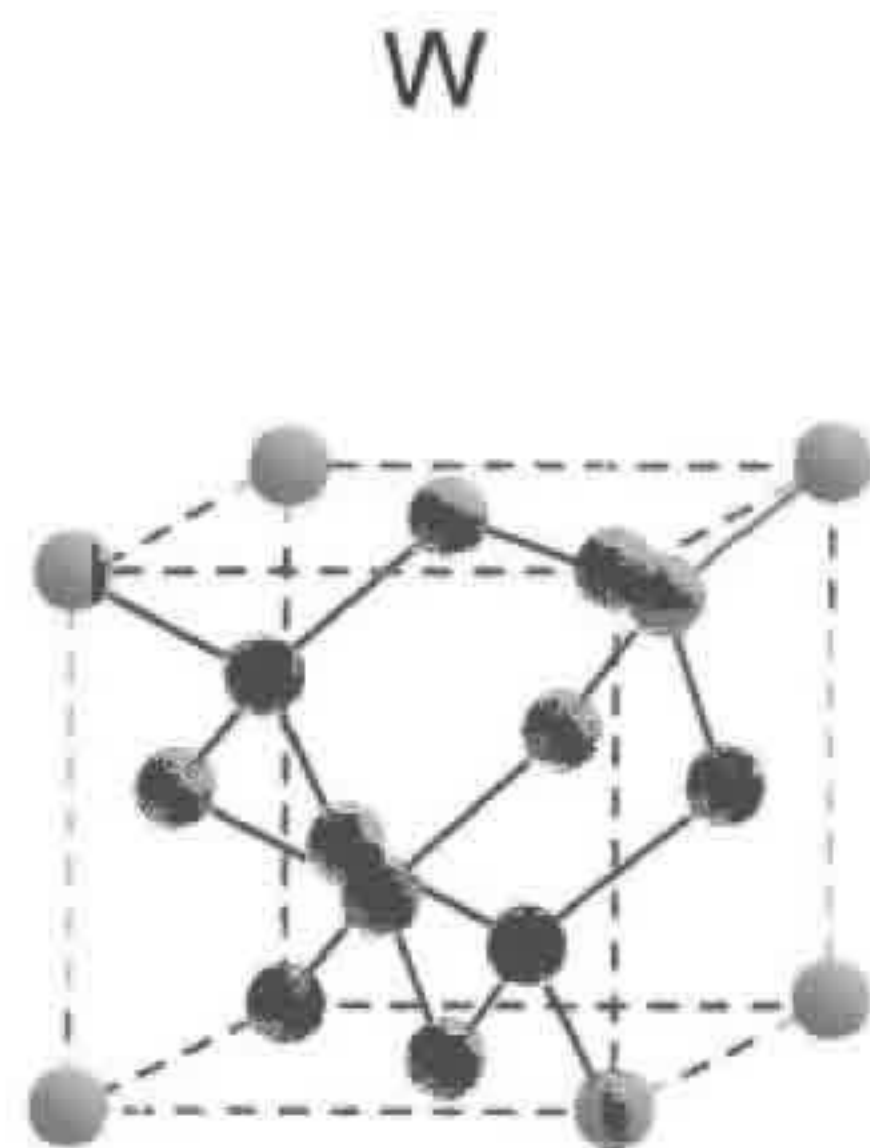
	Electron	Electron
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- C. Pressure increases by 50 %.
- D. Pressure decreases by 50 %.

7. Which row shows the correct formula and electron configuration of the species in calcium nitride?

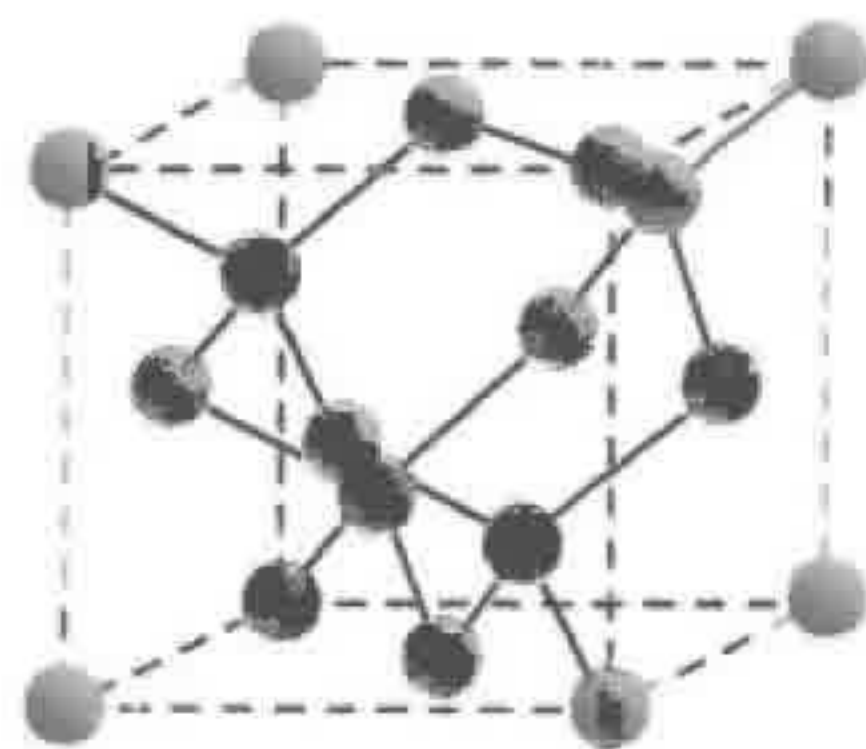
	Formula	Electron configuration of the Ca species	Electron configuration of the N species
A.	Ca_2N_3	[Ar]	[Ne]
B.	Ca_2N_3	[Ar] $4s^2$	[He] $2s^2 2p^3$
C.	Ca_3N_2	[Ar] $4s^2$	[He] $2s^2 2p^3$
D.	Ca_3N_2	[Ar]	[Ne]

8. Diagrams of four stable allotropes of carbon are shown. Which row correctly identifies each structure?

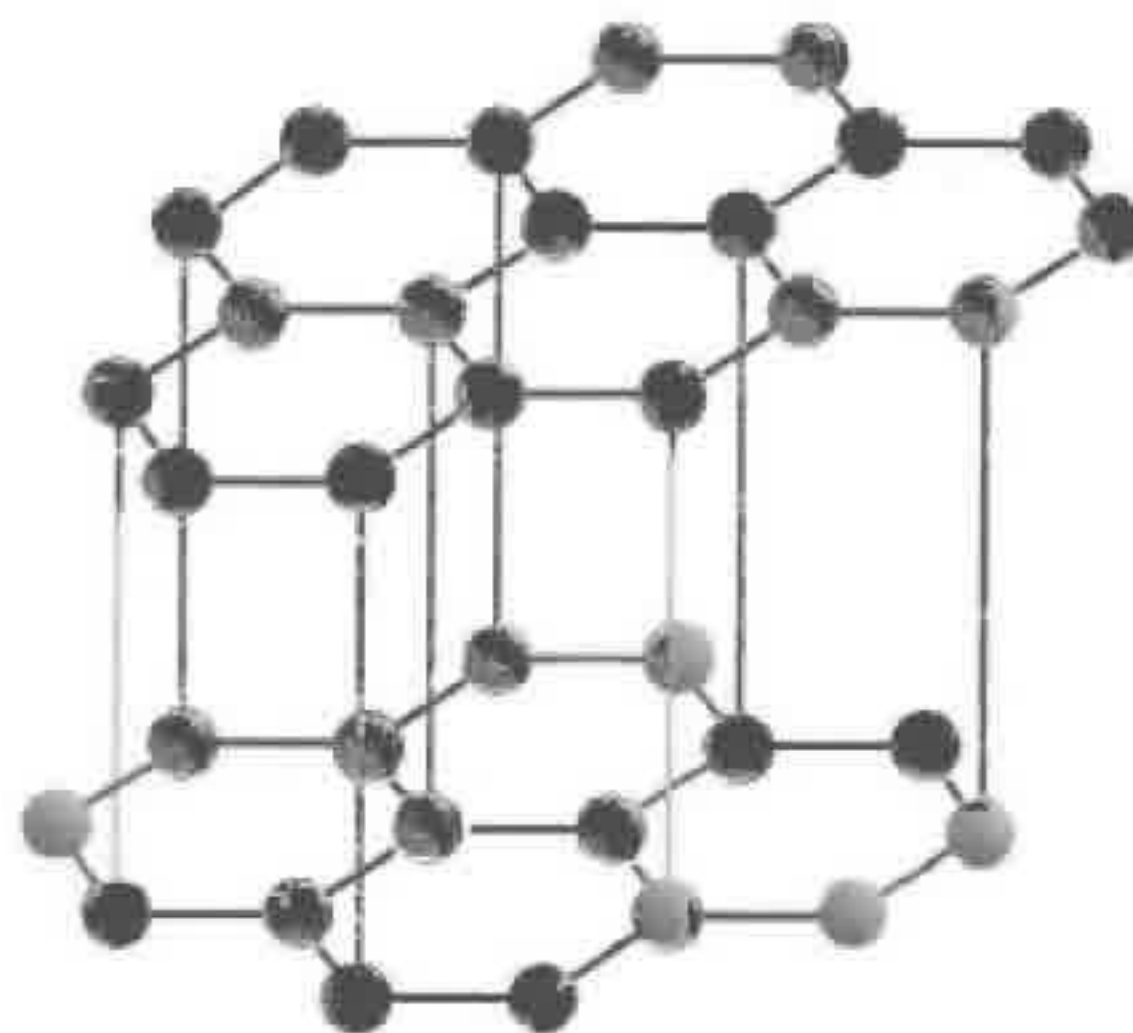


	Graphene	Fullerene	Graphite	Diamond
A.	Z	Y	X	W
B.	W	Z	Y	X

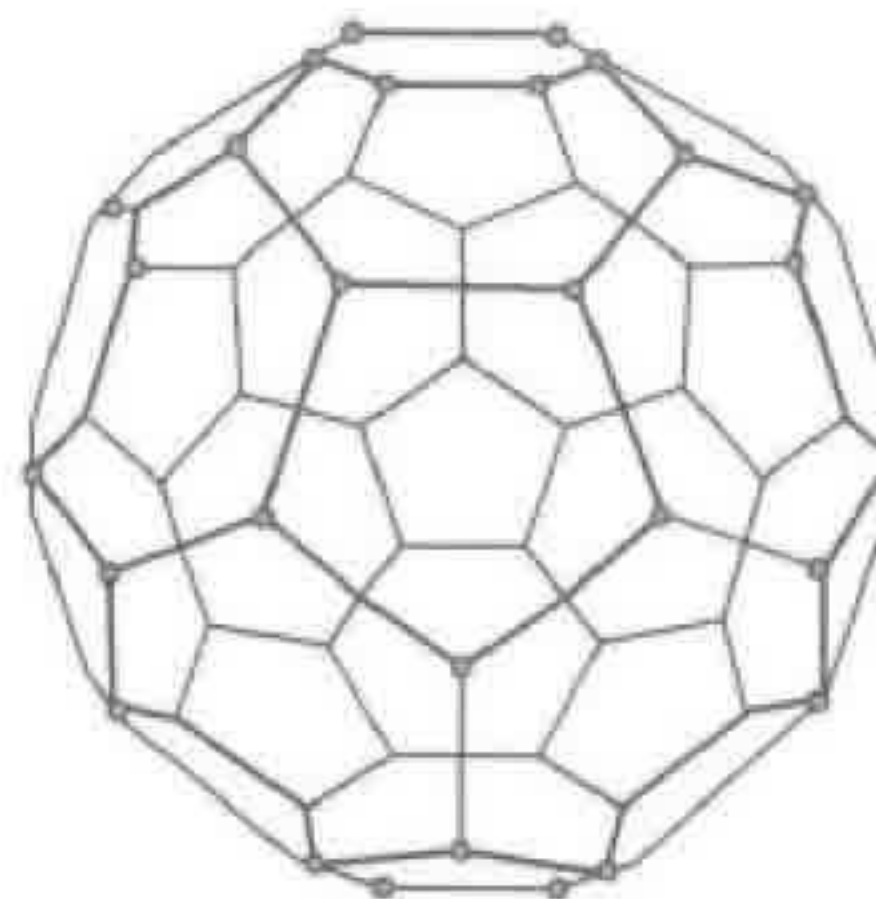
W



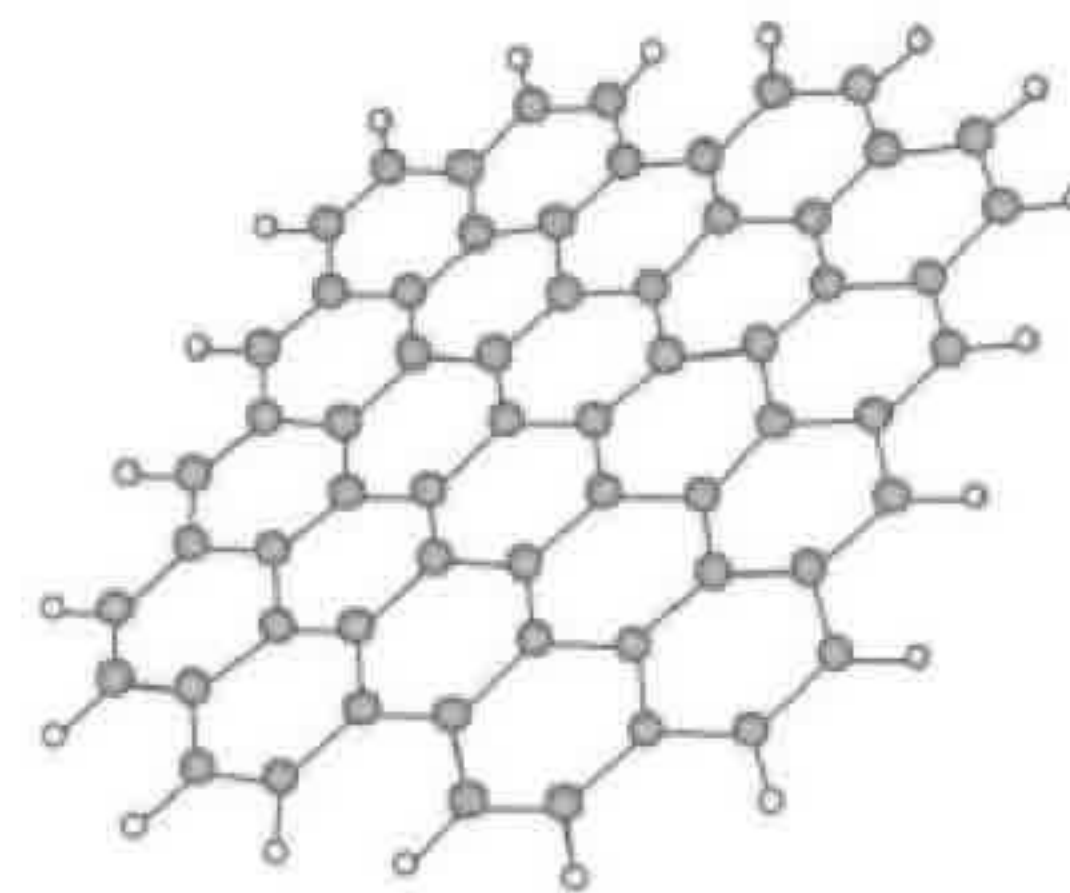
X



Y



Z



A.

Z

Y

X

W

B.

W

Z

Y

X

C.

X

W

Z

Y

D.

Y

X

W

Z

Graphene

Fullerene

Graphite

Diamond

C.	X	W	Z	Y
D.	Y	X	W	Z

9. In which of the following molecules do all atoms obey the octet rule?



A. I and II only

B. I and III only

C. II and III only

D. I, II and III

10. Which compound has the weakest ionic bond?

A. MgO

B. MgS

C. CaO

D. CaS

11. Which row puts these molecules in order of decreasing bond angle?

A. NH_3 , H_2O , C_2H_4 , CH_4

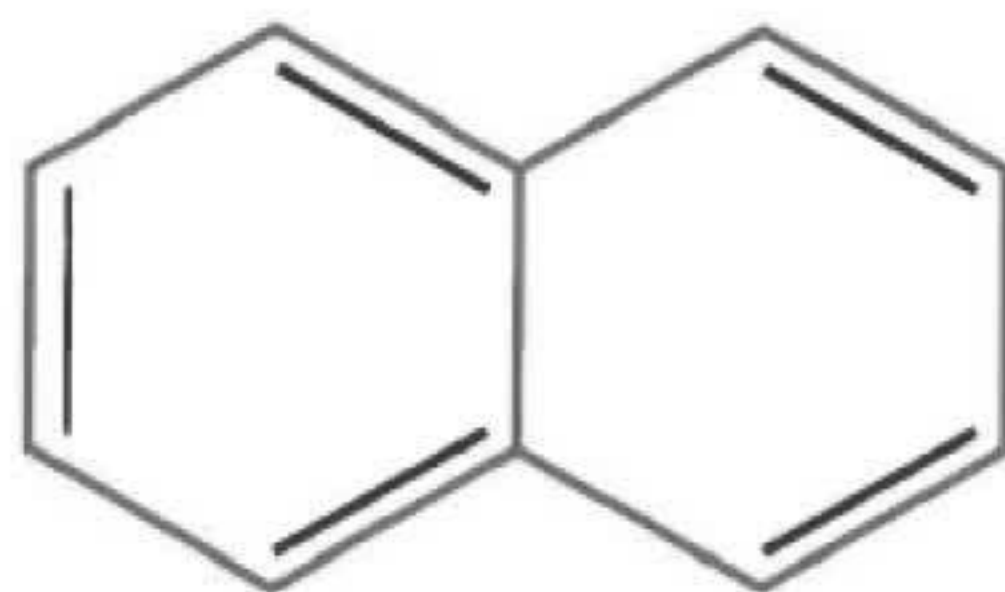
B. H_2O , C_2H_4 , CH_4 , NH_3

C. C_2H_4 , CH_4 , NH_3 , H_2O

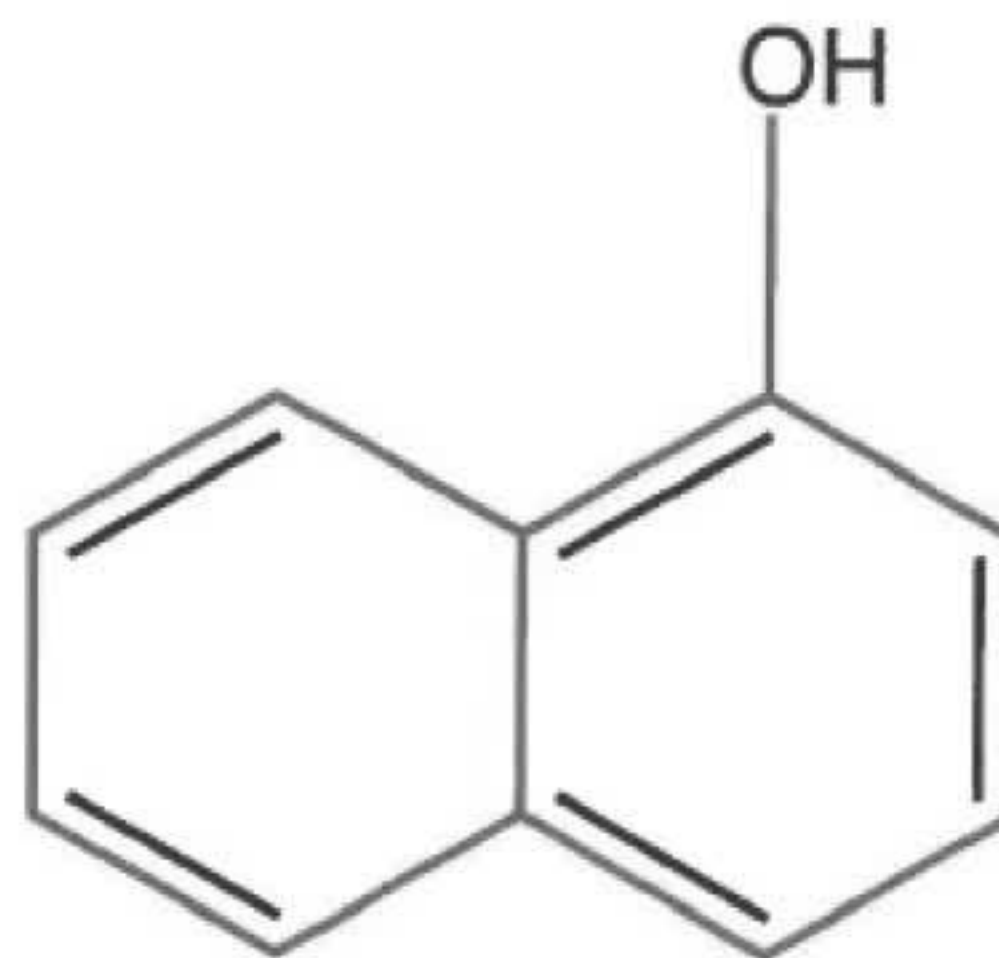
D. CH_4 , NH_3 , H_2O , C_2H_4



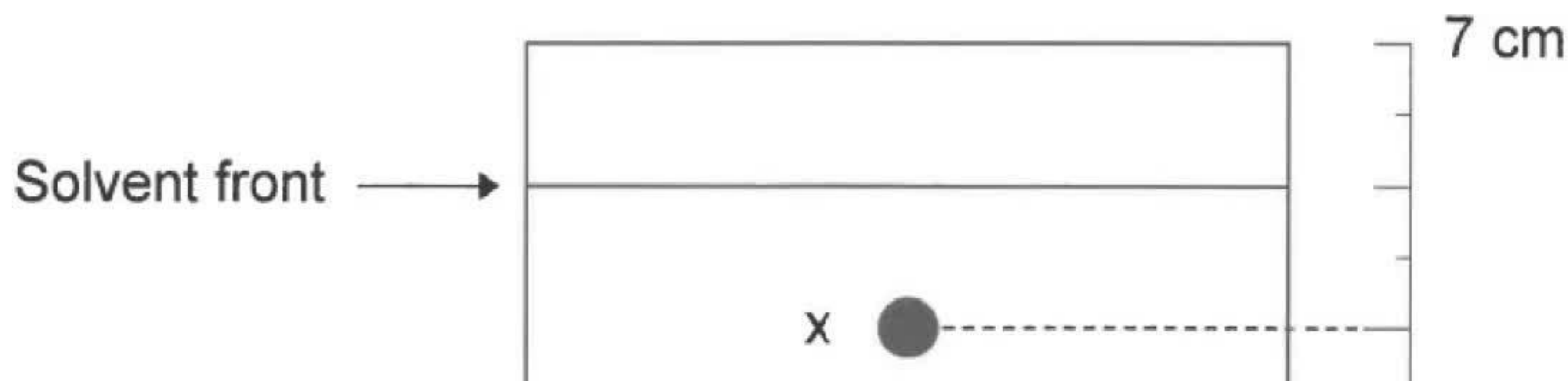
12. The following thin-layer chromatogram was obtained from a mixture of naphthalene, $C_{10}H_8$, and naphthol, $C_{10}H_8O$. A polar silica was used as the stationary phase, and a non-polar hexane as the mobile phase.

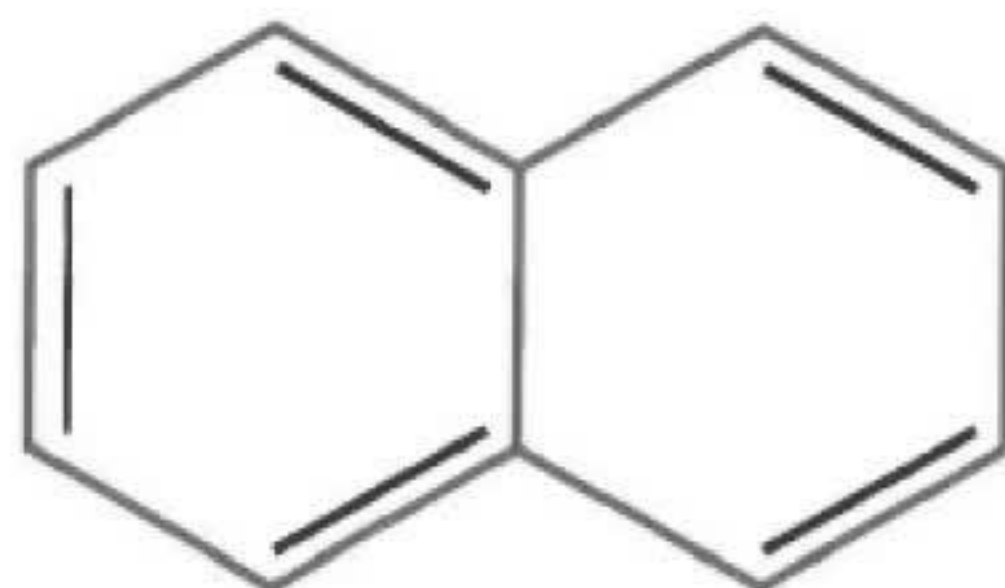


Naphthalene

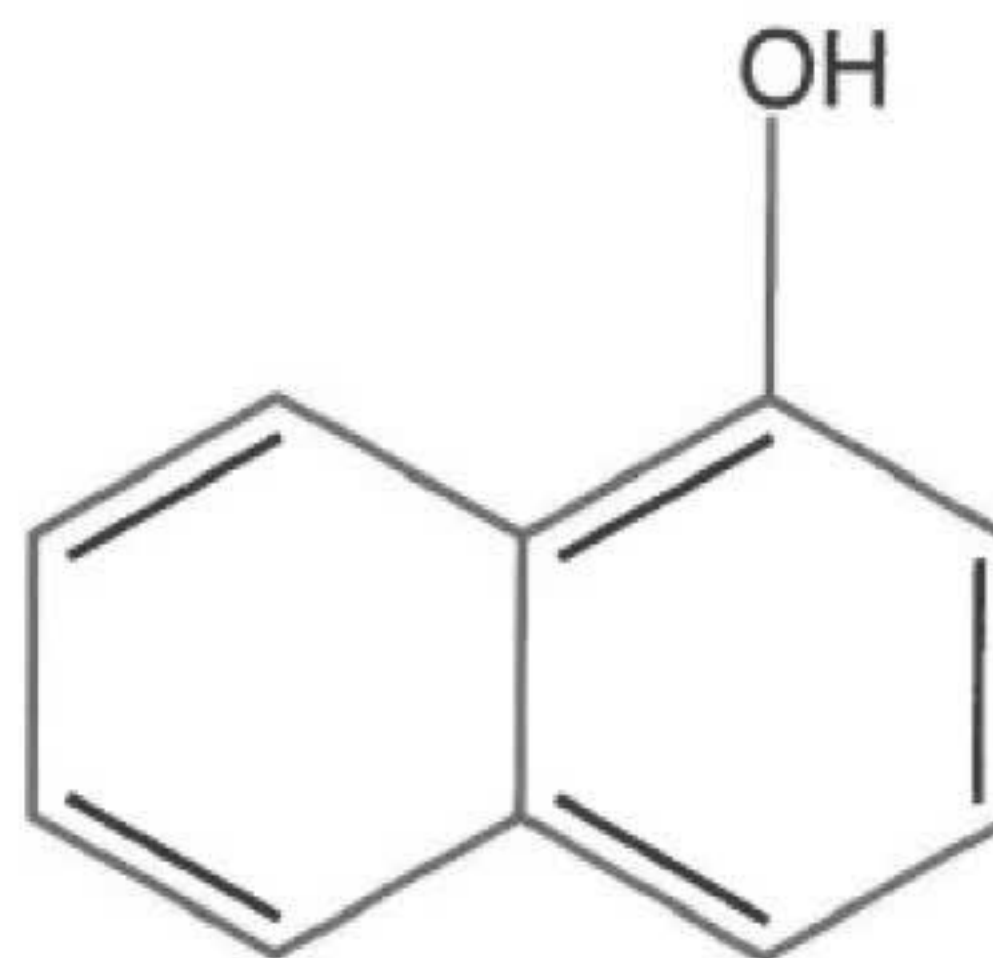


Naphthol

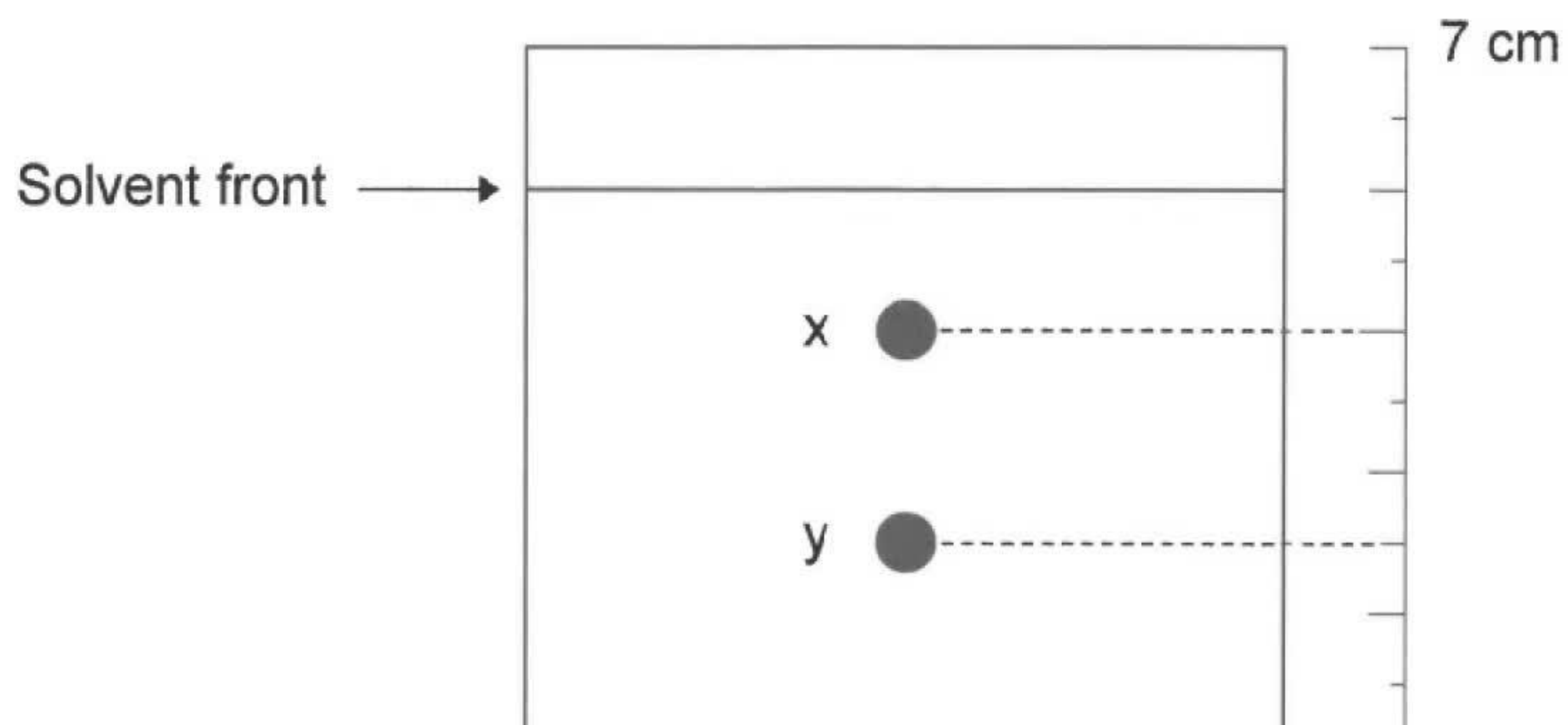


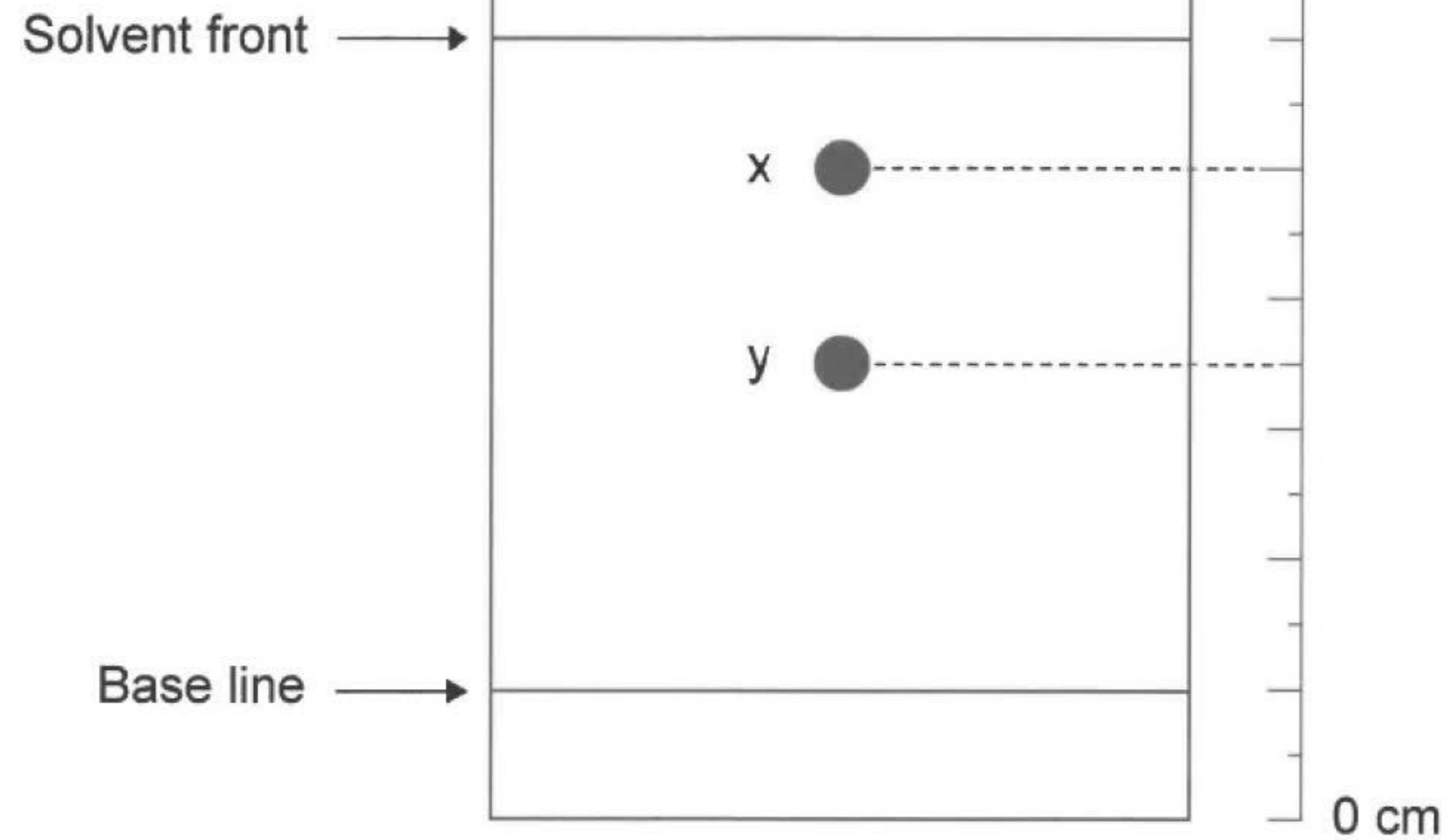


Naphthalene



Naphthol





What are the position and retardation factor (R_F) of naphthalene?

- A. y and 0.50
- B. y and 0.58
- C. x and 0.80
- D. x and 0.83

13. Aluminium trichloride, $\text{AlCl}_3(\text{g})$, and beryllium chloride, $\text{BeCl}_2(\text{g})$, have been known for many years. Which adaptation of the bonding model was made to explain the bonding in $\text{AlCl}_3(\text{g})$ and $\text{BeCl}_2(\text{g})$?

- A. The existence of coordination bonding was proposed.
- B. Ideas about oxidation states were formed.
- C. Exceptions to the octet rule were introduced.
- D. Lewis formulas of molecules were developed.

14. Germanium has several stable oxidation states. Which row shows these species in order of increasing radius?



- A. Ge^{4+} , Ge^{2+} , Ge , Ge^{4-} .
- B. Ge^{4-} , Ge^{4+} , Ge^{2+} , Ge .
- C. Ge , Ge^{4-} , Ge^{4+} , Ge^{2+} .
- D. Ge^{2+} , Ge , Ge^{4-} , Ge^{4+} .

15. Which statement explains the trend in first ionisation energy from sodium Na to chlorine Cl ?

15. Which statement explains the trend in first ionisation energy from sodium, Na, to chlorine, Cl?
- A. Nuclear charge increases.
 - B. Electronegativity decreases.
 - C. Atomic radius increases.
 - D. Shielding decreases.
16. Which elements would react with each other most vigorously?
- A. K(s) and Br₂(g)
 - B. K(s) and Cl₂(g)
 - C. Na(s) and Br₂(g)
 - D. Na(s) and Cl₂(g)



17. Ammonia, NH_3 and nitrous acid, HNO_2 , are compounds of nitrogen. What are the oxidation states of nitrogen in these compounds?

	NH_3	HNO_2
A.	+3	+3
B.	-3	-3
C.	+3	-3
D.	-3	+3

18. Consider a reaction mixture that is in thermal equilibrium with the surroundings. When a reaction takes place, the temperature of the mixture decreases.

Which row correctly shows the changes in energy when the new thermal equilibrium is established?

	Energy of system	Energy of surroundings
A.	decreases	decreases

C.	+3	-3
D.	-3	+3

18. Consider a reaction mixture that is in thermal equilibrium with the surroundings. When a reaction takes place, the temperature of the mixture decreases.

Which row correctly shows the changes in energy when the new thermal equilibrium is established?

	Energy of system	Energy of surroundings
A.	decreases	decreases
B.	increases	increases
C.	decreases	increases
D.	increases	decreases

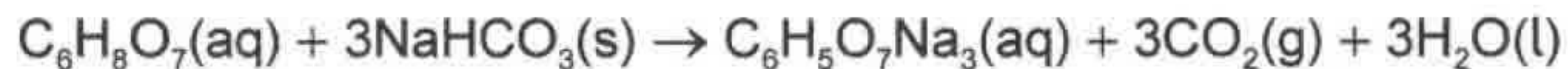
19. 0.050 mol of sodium hydrogencarbonate was added to 50 cm³ of 2.0 mol dm⁻³ solution of citric acid.



The initial temperature of the solution was 298 K and the final temperature was 297 K

C.	decreases	increases
D.	increases	decreases

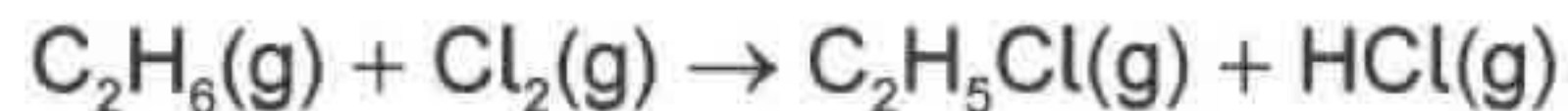
19. 0.050 mol of sodium hydrogencarbonate was added to 50 cm³ of 2.0 mol dm⁻³ solution of citric acid.



The initial temperature of the solution was 298 K and the final temperature was 297 K.
What is the enthalpy of this reaction in kJ mol⁻¹? ($c_w = 4.18 \text{ J g}^{-1} \text{ K}^{-1}$, $Q = mc\Delta T$)

- A. +4.18
- B. -4.18
- C. +12.5
- D. -12.5

20. Chloroethane is produced by the reaction of ethane with chlorine.

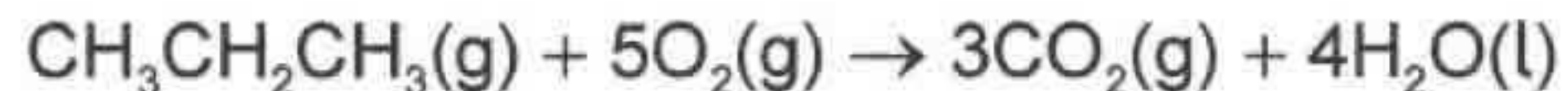


($\text{C}_2\text{H}_6 = 30.08 \text{ g mol}^{-1}$, $\text{Cl}_2 = 70.90 \text{ g mol}^{-1}$, $\text{C}_2\text{H}_5\text{Cl} = 64.52 \text{ g mol}^{-1}$, $\text{HCl} = 36.46 \text{ g mol}^{-1}$)

$$\% \text{ atom economy} = \frac{\text{molar mass of desired product}}{\text{molar mass of all reactants}} \times 100$$

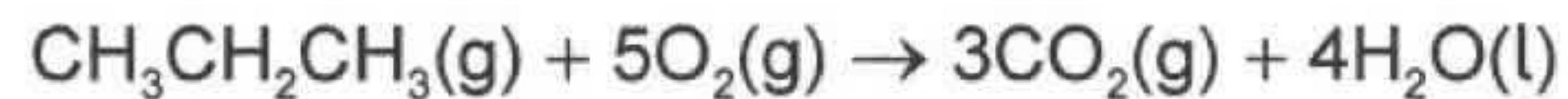
What is the atom economy for this reaction?

- A. 15.7 %
 - B. 46.6 %
 - C. 63.9 %
 - D. 99.9 %
21. 75 cm^3 of propane was completely combusted in 400 cm^3 of oxygen according to the following equation.



D. 99.9 %

21. 75 cm³ of propane was completely combusted in 400 cm³ of oxygen according to the following equation.



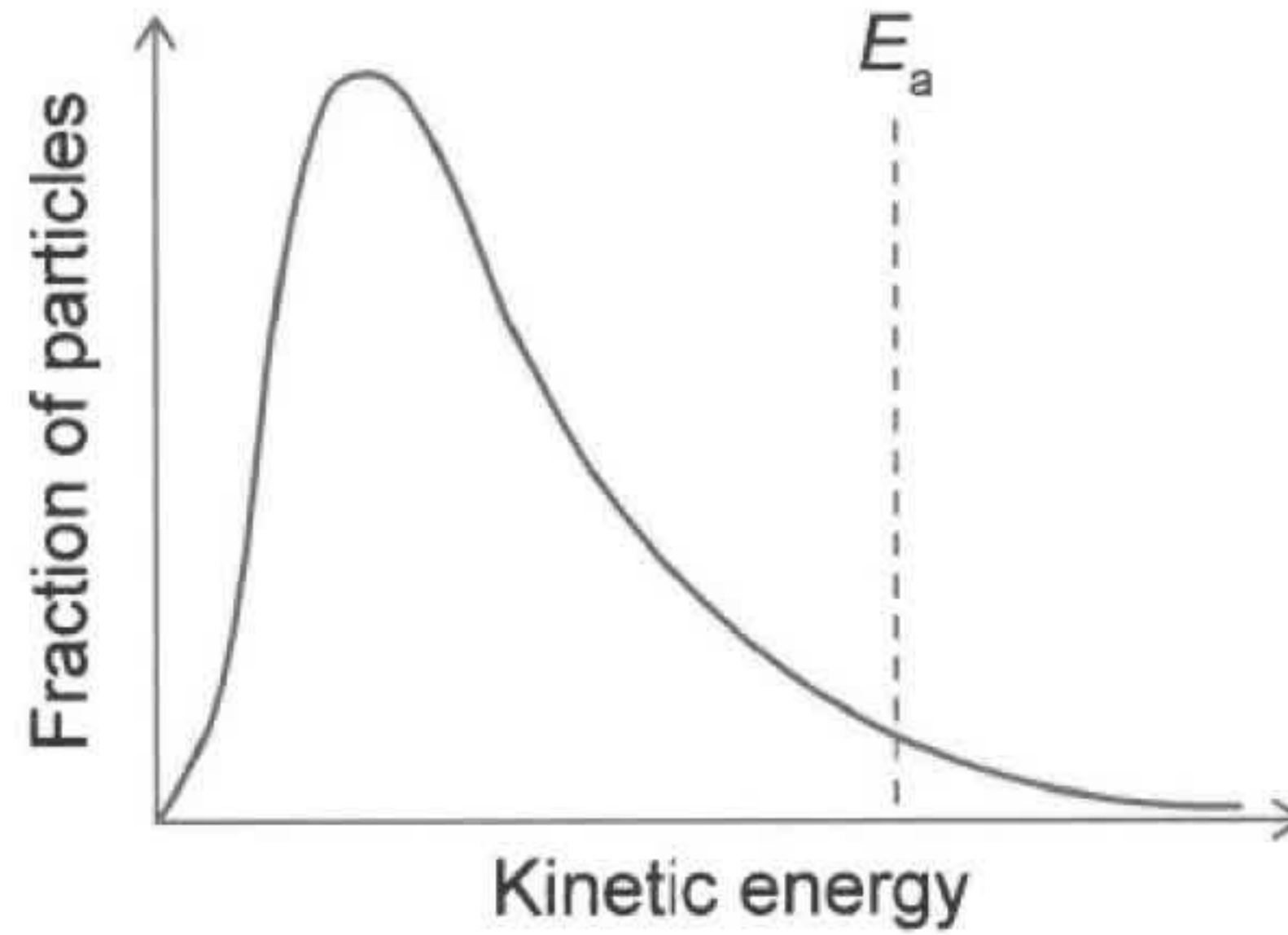
What is the volume of unreacted oxygen remaining at the original conditions?

- A. 25 cm³
- B. 50 cm³
- C. 75 cm³
- D. 150 cm³

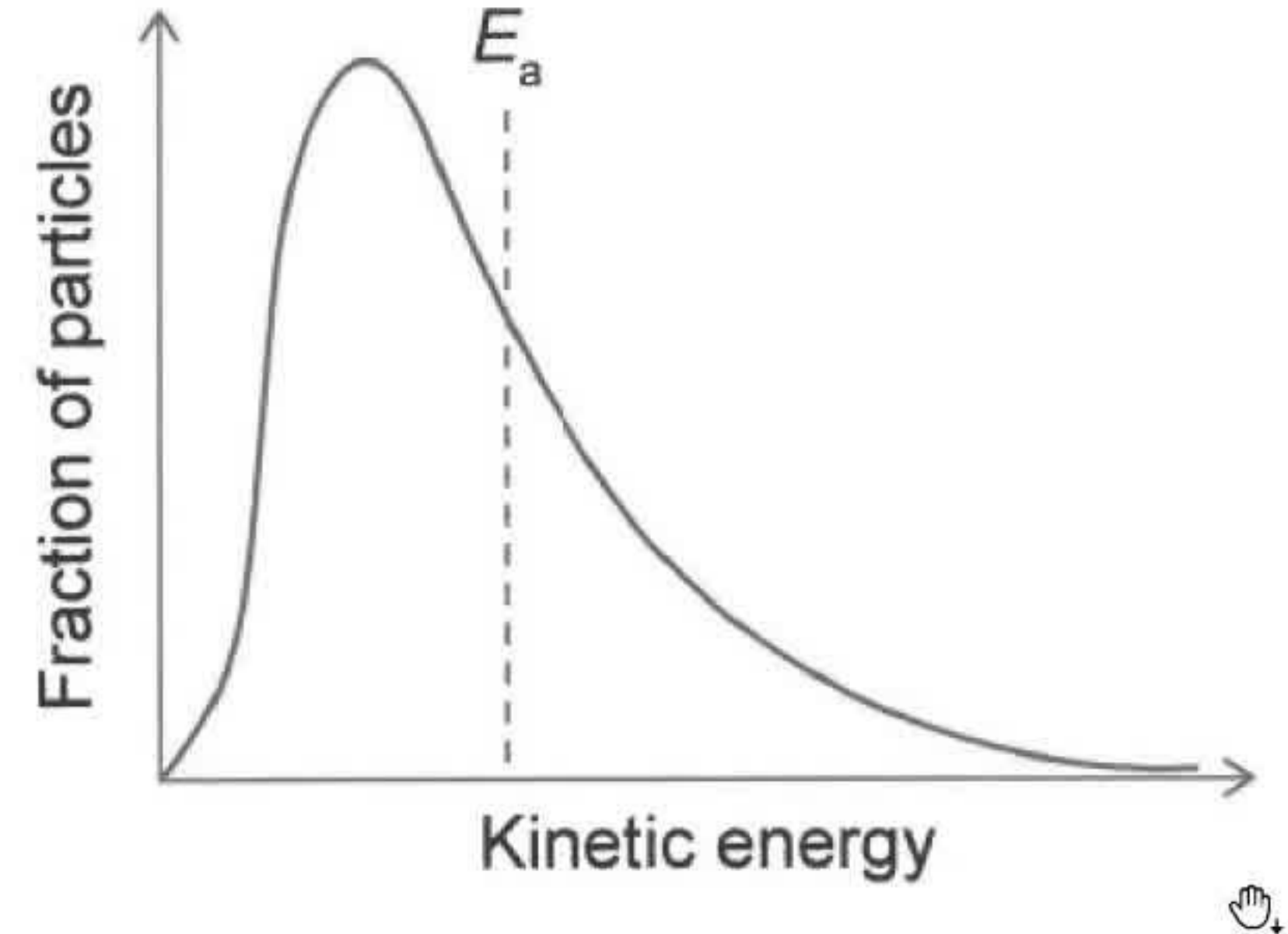
22. Consider the following Maxwell–Boltzmann energy distribution at a certain temperature.

Which represents the distribution and value of activation energy at a lower temperature?

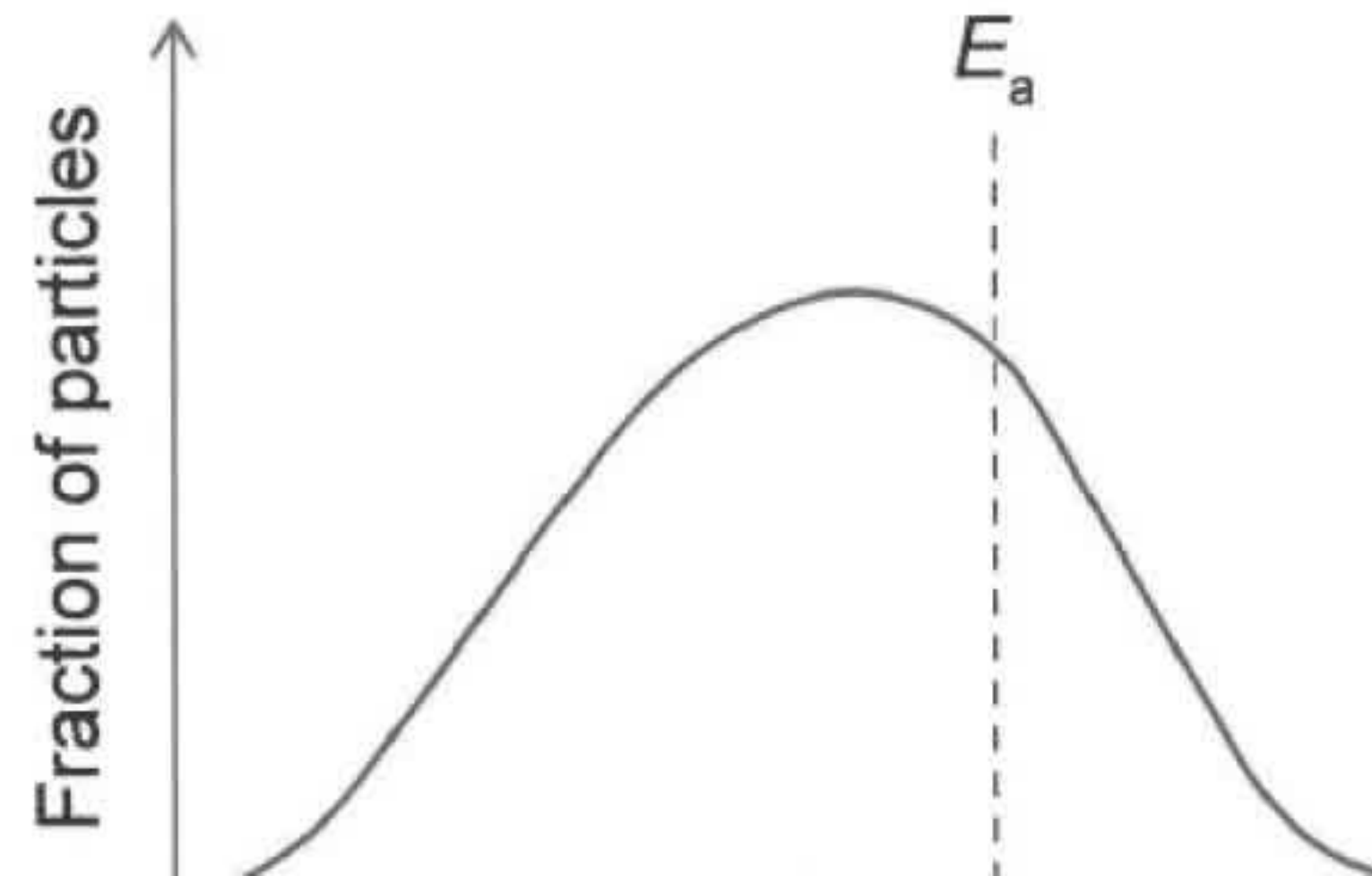
A.



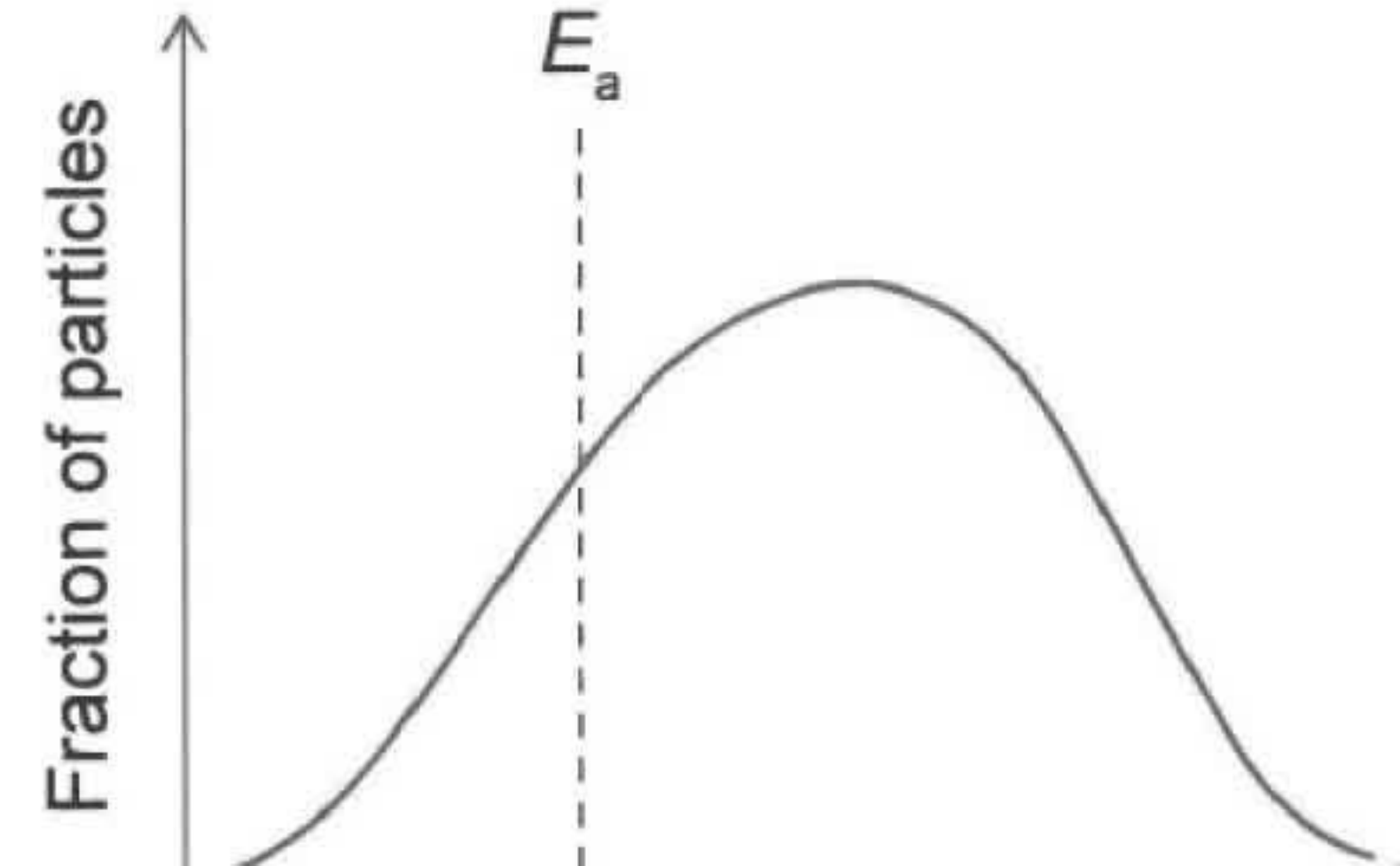
B.



C.

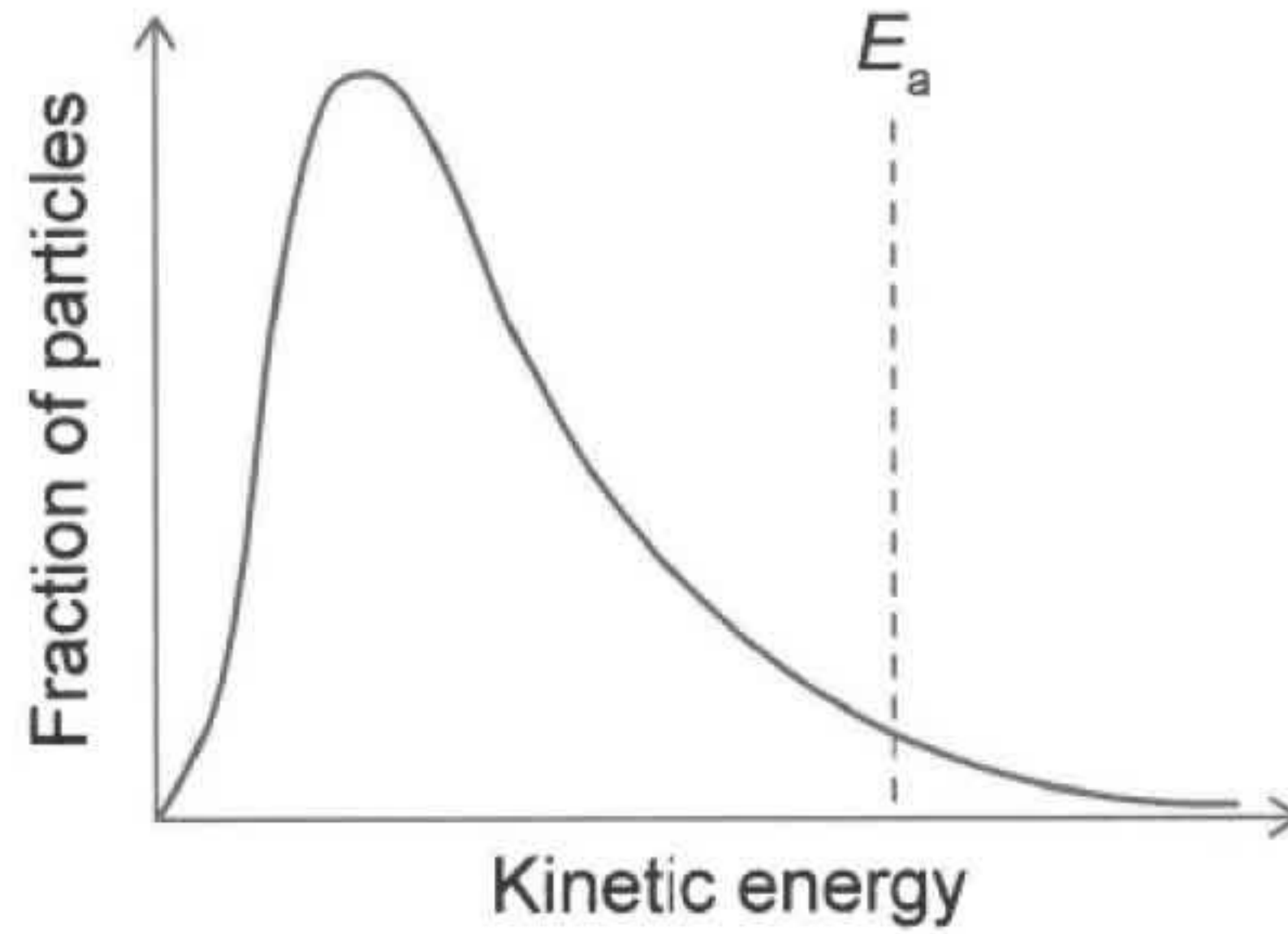


D.

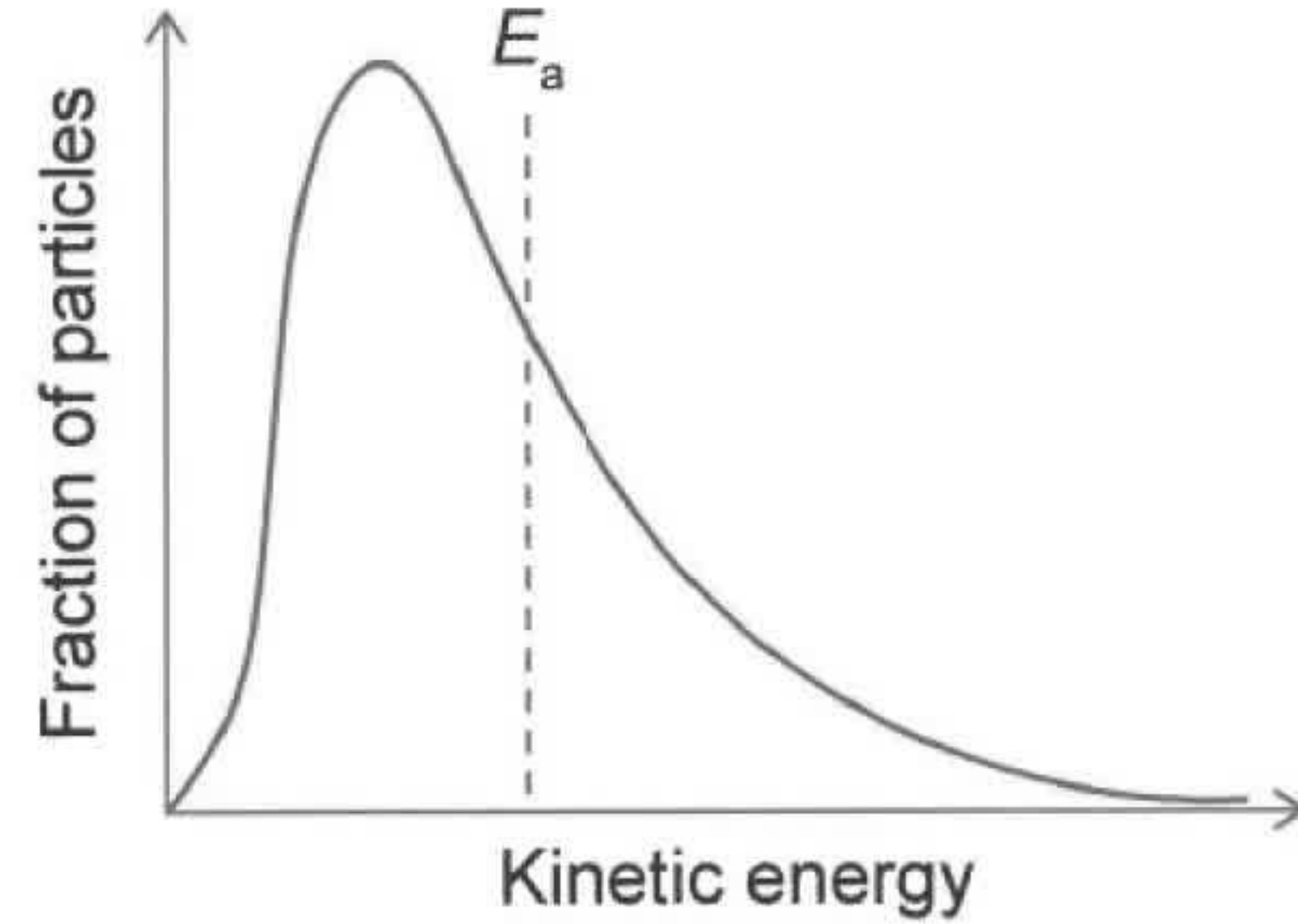


Which represents the distribution and value of activation energy at a lower temperature?

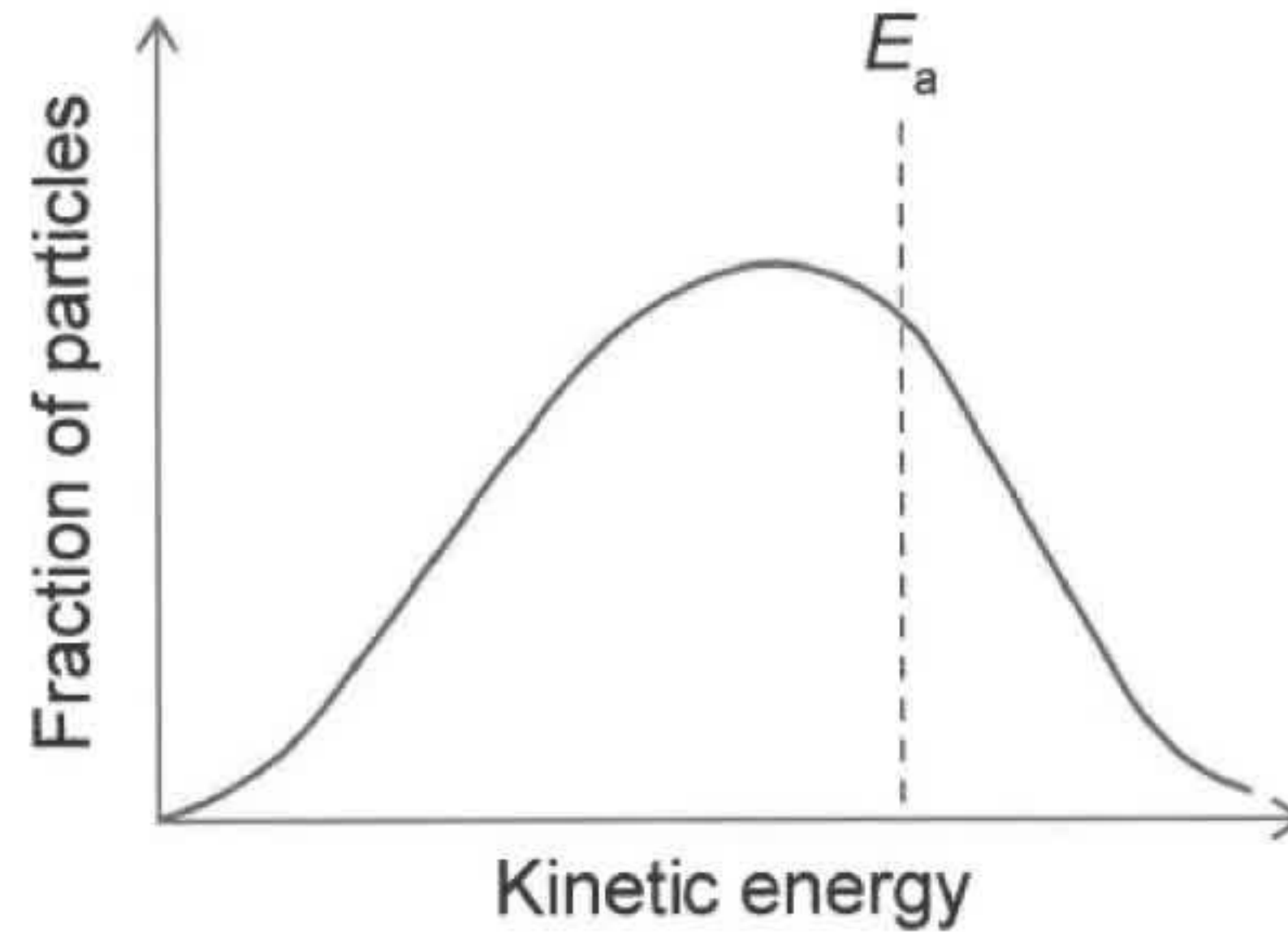
A.



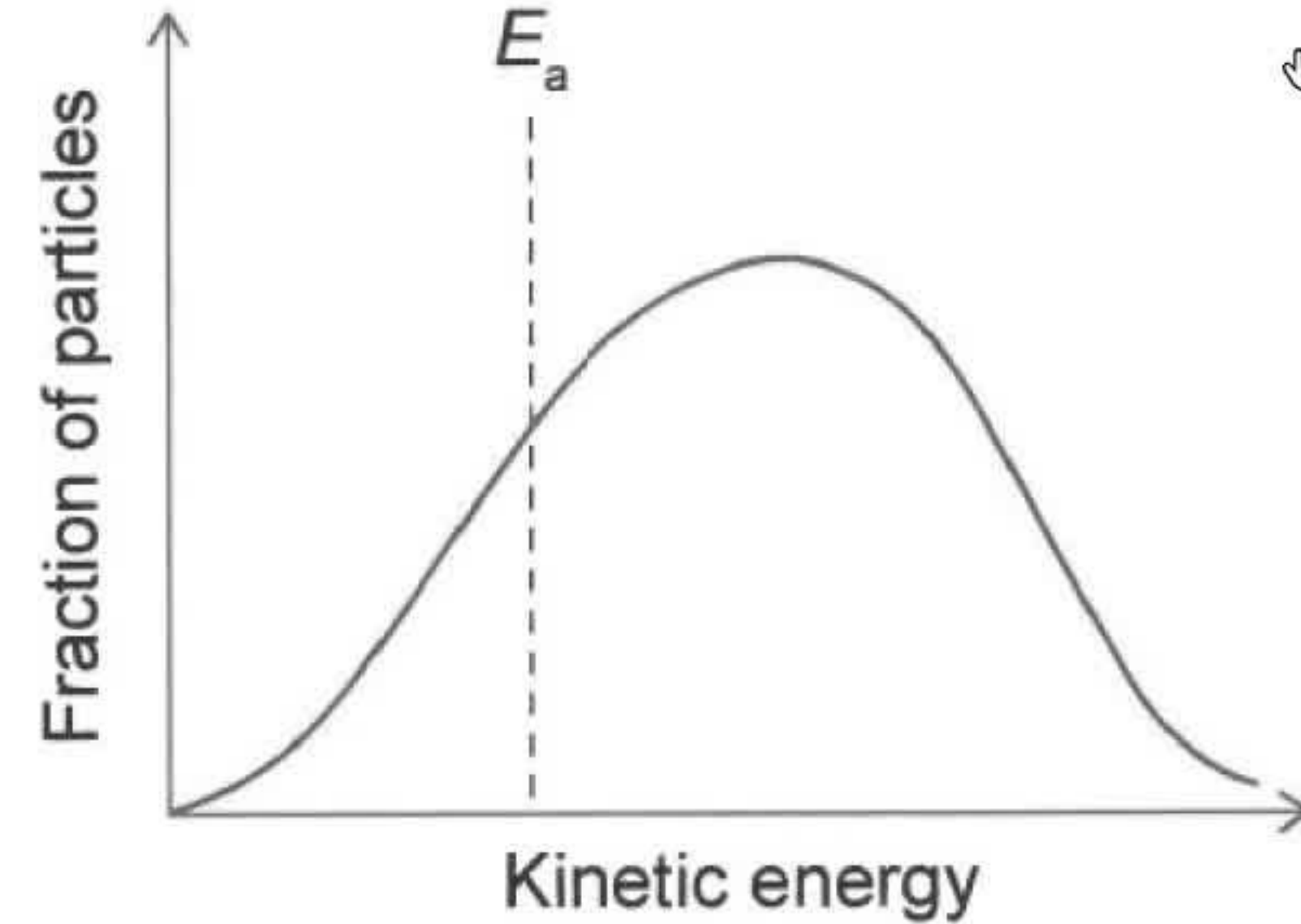
B.



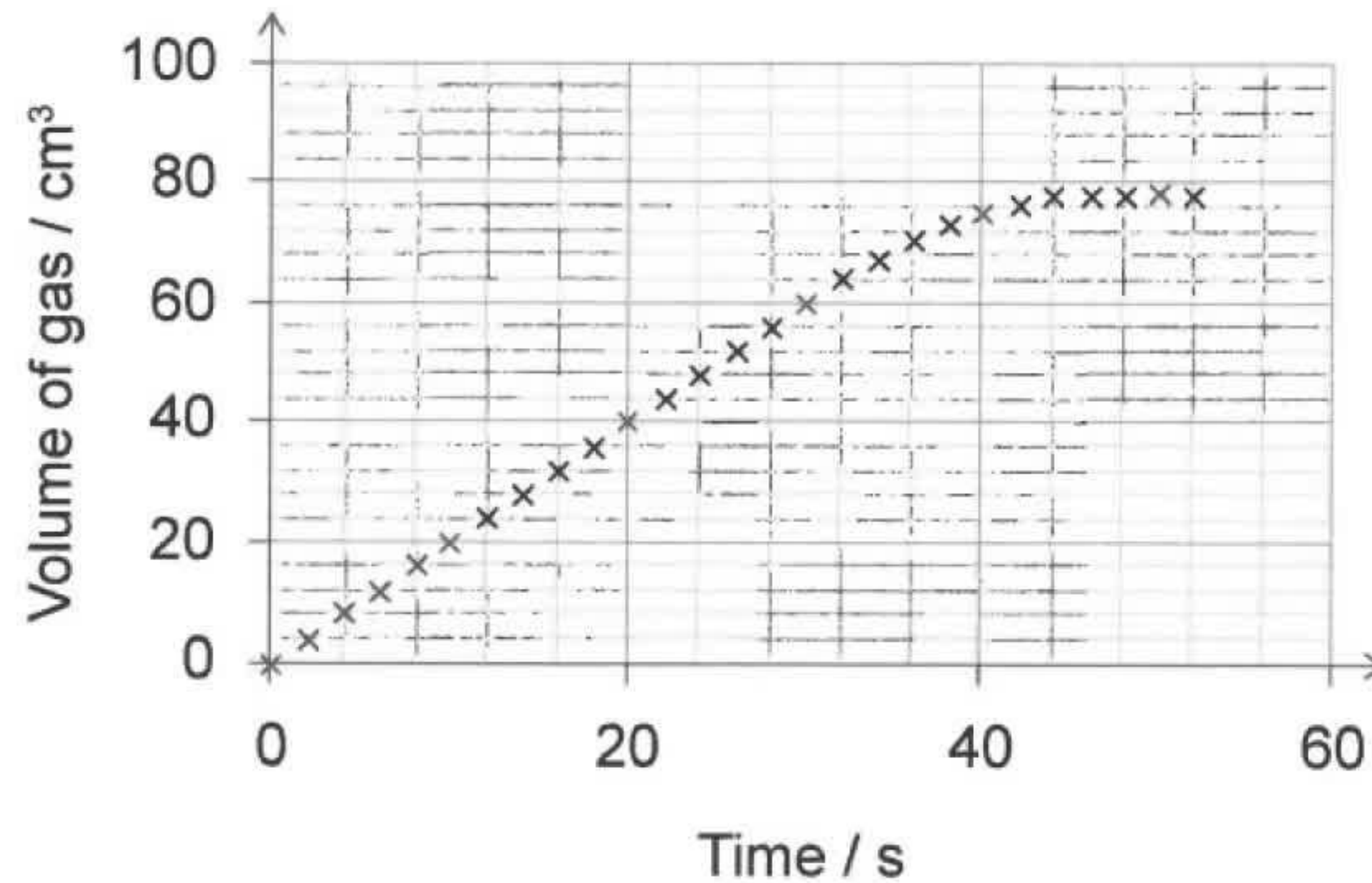
C.



D.



23. The volume of gas produced in a kinetics experiment is plotted against time.



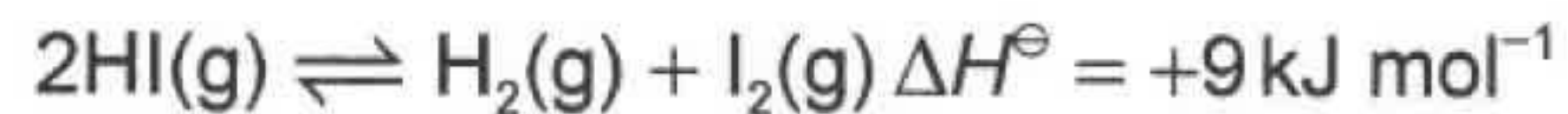
What is the average rate of gas production for the first 20 seconds?

- A. $2.0 \text{ cm}^3 \text{ s}^{-1}$
- B. $1.5 \text{ cm}^3 \text{ s}^{-1}$
- C. $1.0 \text{ cm}^3 \text{ s}^{-1}$
- D. $0.5 \text{ cm}^3 \text{ s}^{-1}$

C. $1.0 \text{ cm}^3 \text{ s}^{-1}$

D. $0.5 \text{ cm}^3 \text{ s}^{-1}$

24. The following equilibrium was established.



Which changes would shift the position of equilibrium furthest to the left?

	Pressure	Temperature
A.	decreases	no change
B.	increases	decreases
C.	decreases	increases
D.	increases	no change

25. Which two species can be both Brønsted–Lowry acids and Brønsted–Lowry bases?
- A. H_3PO_4 and PO_4^{3-}
 - B. HPO_4^{2-} and H_2PO_4^-
 - C. HPO_4^{2-} and PO_4^{3-}
 - D. PO_4^{3-} and H_2PO_4^-
26. A solution has a pH of 3.0. What is the hydrogen ion concentration in the solution in mol dm^{-3} ?
- A. 3.0×10^{-3}
 - B. 1.0×10^{-3}
 - C. 1.0×10^3
 - D. 3.0×10^3
27. A reactivity series of selected elements, arranged from highest activity to lowest, is shown.

27. A reactivity series of selected elements, arranged from highest activity to lowest, is shown.

Zn
Fe
Cd
Co
Ni
Sn
Pb
Cu

Which reaction is spontaneous?

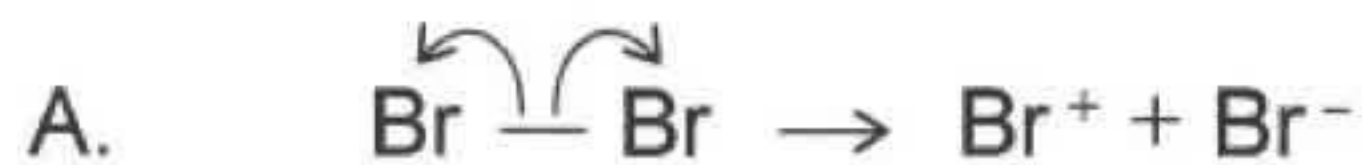
- A. $\text{Cd}^{2+}(\text{aq}) + \text{Pb}(\text{s}) \rightarrow \text{Pb}^{2+}(\text{aq}) + \text{Cd}(\text{s})$
- B. $\text{Fe}^{2+}(\text{aq}) + \text{Sn}(\text{s}) \rightarrow \text{Sn}^{2+}(\text{aq}) + \text{Fe}(\text{s})$
- C. $\text{Zn}^{2+}(\text{aq}) + \text{Cu}(\text{s}) \rightarrow \text{Cu}^{2+}(\text{aq}) + \text{Zn}(\text{s})$
- D. $\text{Ni}^{2+}(\text{aq}) + \text{Co}(\text{s}) \rightarrow \text{Co}^{2+}(\text{aq}) + \text{Ni}(\text{s})$

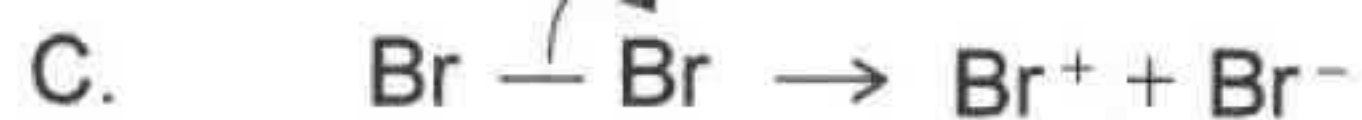


28. Which alcohol and conditions would produce the highest yield of methylpropanoic acid?

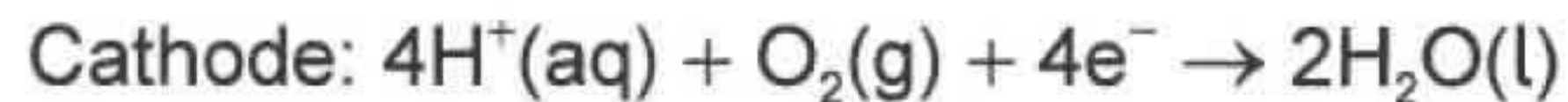
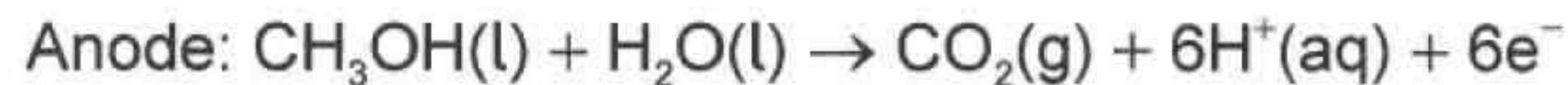
	Alcohol	Conditions
A.	2-methylpropan-2-ol	Reflux
B.	2-methylpropan-1-ol	Reflux
C.	2-methylpropan-2-ol	Distillation
D.	2-methylpropan-1-ol	Distillation

29. Which row correctly shows heterolytic fission?





30. The methanol fuel cell relies on the oxidation of methanol by oxygen. The two half-equations are as follows:



What is the balanced equation for the overall reaction?

- A. $2\text{CH}_3\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{l})$
- B. $\text{CH}_3\text{OH}(\text{l}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$
- C. $2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{CH}_3\text{OH}(\text{l}) + 3\text{O}_2(\text{g})$
- D. $\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow \text{CH}_3\text{OH}(\text{l}) + 2\text{O}_2(\text{g})$
-