

APOLLO

Electron sharing reactions

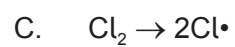
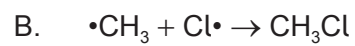
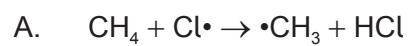
SL

Chemistry

First assessment 2025

4 questions | questions

30. Which equation represents propagation in a radical substitution mechanism?

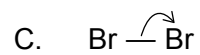
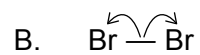


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29. Which radical is most likely to form during the breakdown of one covalent bond of dichlorofluoromethane, CHCl_2F , in the upper atmosphere?

- A. $\bullet\text{CHClF}$
- B. $\bullet\text{CCl}_2\text{F}$
- C. $\bullet\text{CHCl}_2$
- D. $\bullet\text{CHCl}_2\text{F}$

29. Which species represents the product of the homolytic fission of Br_2 in the presence of UV light?



29. Which statements explain the following reactions occurring in the upper atmosphere?

	Chlorofluorocarbon (CFC) compounds break down to produce chlorine radicals but usually not fluorine radicals.	A single chlorine radical breaks down many ozone, O₃, molecules.
A.	C–Cl bond is stronger than C–F bond	chain propagation steps produce more radicals
B.	C–F bond is stronger than C–Cl bond	chain termination steps cause chlorine radicals to reform chlorine molecules
C.	C–Cl bond is stronger than C–F bond	chain termination steps cause chlorine radicals to reform chlorine molecules
D.	C–F bond is stronger than C–Cl bond	chain propagation steps produce more radicals